

RADIOACTIVE DECAY

Radioactive Decay of Candium 150

Introduction: Some isotopes spontaneously emit particles or energy to reach a different, more stable, state. This decay occurs naturally and is not affected by changes in temperature or pressure.

Radioactive decay takes place in a random way. Although you cannot predict just when a given atom will decay, you can predict that all atoms of a given isotope will decay at the same rate.

Radioactive decay is given number of decay of a substance and predictable way.

Half-life of radioactive substance is the time taken for the substance to decay to half of its original amount. It is a constant value for a given substance.

Comparing the percentage of original amount present at any time to the percentage of the original amount present at the start of the half-life, the half-life of the substance can be calculated. The half-life of the substance is a constant value for a given substance.

Procedure:

- 1. Prepare the substance.
- 2. Measure the initial amount.
- 3. Measure the amount after a fixed time interval.
- 4. Calculate the percentage of the original amount remaining.
- 5. Repeat steps 3 and 4 for several more time intervals.

4. Count and record the number of undecayed "undecayed" Candium atoms (150) and the number of decayed atoms.

Repeat steps 3, 4 and 5 until all the Candium atoms have decayed (150) and the number of decayed atoms is 150.

Half-life (in 15 seconds)	"Parent" atoms (all of "undecayed" radioactive Candium atoms remaining after the "half-life" interval)	"Daughter" atoms (all of decayed atoms)
0	150	0
1		
2		
3		
4		
5		
6		
7		
8		
9		
10		

Graph: Use the graph on the following page to represent the data collected throughout the procedure. The half-life, daughter atoms (green) and parent atoms (red) should be plotted on your graph.

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