#### General Chemistry II Chapter 15 Study Guide

- the total energy of the universe remains constant
  - $-\Delta E_{min} = \Delta E_{min} + \Delta E_{marr} = 0$
  - ΔH is often used to approximate ΔE; for enthalpy there is no zero point, we can only measure changes in enthalpy
- The total entropy of the universe increases
  - $\Delta S_{\text{matr}} = \Delta S_{\text{res}} + \Delta S_{\text{matr}} > 0$
  - For entropy there is a zero point, and we can determine the absolute entropy values

#### Third law of thermodynamics:

- Third law of thermodynamics: the entropy of a perfect crystalline substance at absolute zero (0 K) is 0
  - S = 0 at 0 K
- A "perfect" crystal has flawless alignment of all its particles; at absolute 0, the particles have minimum energy, so there is only one microstate

#### Entropy change in phase changes:

- Isothermal process (ex. Phase change): the change in entropy is equal to the heat that would be transferred, divided by the temperature:
  - $\Delta S = \Delta H / T$
- Consider the melting of 1 mol of an ice cube (molar enthalpy change of ice melting= 6.01x10<sup>3</sup> J/mol) in the palm of your hand, determine the ΔS for the ice cube and for its surroundings:

$$\Delta S_{sys} = \frac{(1mol)(6.01x10^{3} J/mol)}{273 K} = 22.0 \frac{J}{K}$$

$$\Delta S_{surr} = \frac{(1mol)(-6.01x10^{3} J/mol)}{310 K} = -19.4 \frac{J}{K}$$

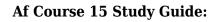
$$\Delta S_{solol} = \Delta S_{sys} + \Delta S_{surr} = 22.0 - 19.4 = 2.6 \frac{J}{K}$$

#### Entropy change in isothermal expansion/compression:

- Entropy change for isothermal expansion/compression of a gas is:
  - $\Delta S_{\text{sys}} = nR \ln V_2 / V_1$ 
    - n= moles of gas
    - R= 8.314 J/molK
    - V<sub>1</sub> = initial volume
    - V<sub>2</sub> = final volume
  - If volume is increased, ΔS<sub>pm</sub> > 0

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