

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

- **Crystalline Solids:** Atoms, ions, or molecules are arranged in a regular, repeating three-dimensional pattern. They have sharp melting points and distinct cleavage planes.
- **Amorphous Solids:** Atoms, ions, or molecules are arranged in a disordered, non-repeating pattern. They do not have sharp melting points and do not cleave along specific planes.
- **Intermolecular Forces:** The strength of these forces determines the melting point and other physical properties of a solid.
 - **London Dispersion Forces:** Present in all solids, they are the weakest intermolecular forces.
 - **Dipole-Dipole Interactions:** Present in polar solids, they are stronger than London forces.
 - **Hydrogen Bonding:** Present in solids containing N-H, O-H, or F-H bonds, they are the strongest intermolecular forces.
- **Unit Cells:** The smallest repeating unit of a crystal lattice. Common types include simple cubic, body-centered cubic (BCC), and face-centered cubic (FCC).
- **Packing Efficiency:** The fraction of space occupied by atoms in a unit cell. FCC and hexagonal close packing (HCP) have the highest efficiency (74%).

Types of Chemical Calculations

- **Stoichiometry:** Calculations involving the relative amounts and masses of substances in a chemical reaction.
- **Limiting Reactant:** The reactant that is completely consumed first, limiting the amount of product that can be formed.
- **Theoretical Yield:** The maximum amount of product that can be formed from a given amount of reactants.
- **Actual Yield:** The amount of product actually obtained from a reaction.
- **Percent Yield:** The ratio of actual yield to theoretical yield, expressed as a percentage.
- **Concentration:** Calculations involving the amount of solute in a given volume of solution (molarity, molality, etc.).
- **Gas Laws:** Calculations involving the pressure, volume, and temperature of a gas.

Periodic Table & Trends

Diagram of the periodic table showing trends for atomic radius, ionization energy, and electronegativity. Arrows indicate the direction of increasing values for each property.

- **Atomic Radius:** Increases from top-left to bottom-right.
- **Ionization Energy:** Increases from bottom-right to top-left.
- **Electronegativity:** Increases from bottom-right to top-left.

Electronegativity Table

- **Electronegativity:** A measure of an atom's ability to attract and hold onto electrons in a chemical bond.
- **Trends:** Electronegativity increases from bottom-left to top-right.
- **Pauling Scale:** The most common scale for electronegativity, ranging from 0.7 to 4.0.

Diagram of the periodic table with electronegativity values for each element. The values range from 0.7 for Francium to 4.0 for Fluorine.

Diagram showing molecular geometries for various central atom configurations. Examples include linear, trigonal planar, tetrahedral, trigonal bipyramidal, and octahedral.

- **Molecular Geometry:** The three-dimensional arrangement of atoms in a molecule.
- **VSEPR Theory:** Valence Shell Electron Pair Repulsion theory, used to predict molecular geometry.
- **Bond Angles:** The angle between two bonds originating from the same central atom.

Intermolecular Forces

- **London Dispersion Forces:** Weak, temporary forces caused by fluctuations in electron density.
- **Dipole-Dipole Interactions:** Forces between polar molecules.
- **Hydrogen Bonding:** Strong intermolecular forces between molecules containing N-H, O-H, or F-H bonds.

Table summarizing the types of intermolecular forces present in different classes of compounds.

Thermodynamics

- **First Law of Thermodynamics:** Energy is conserved; it can be converted from one form to another but cannot be created or destroyed.
- **Second Law of Thermodynamics:** The entropy of the universe always increases.
- **Third Law of Thermodynamics:** The entropy of a perfect crystal is zero at absolute zero.

Equilibrium

- **Chemical Equilibrium:** A state where the concentrations of reactants and products remain constant over time.
- **Equilibrium Constant (K):** A ratio of product concentrations to reactant concentrations, each raised to the power of its stoichiometric coefficient.
- **Le Chatelier's Principle:** A system at equilibrium will shift to counteract any stress applied to it.

Acids & Bases

- **Arrhenius Definition:** Acids produce H^+ ions in water, and bases produce OH^- ions.
- **Bronsted-Lowry Definition:** Acids are proton donors, and bases are proton acceptors.
- **Conjugate Acid-Base Pairs:** Two species that differ by one proton.

Redox

- **Oxidation:** Loss of electrons or increase in oxidation state.
- **Reduction:** Gain of electrons or decrease in oxidation state.
- **Oxidation States:** A number assigned to an element in a compound representing its degree of oxidation.

Electrochemistry

- **Galvanic Cells:** Devices that convert chemical energy into electrical energy.
- **Standard Reduction Potentials:** A measure of the tendency of a species to be reduced.
- **Nernst Equation:** Relates the cell potential to the concentrations of the reactants and products.

Chemical Kinetics

- **Reaction Rate:** The change in concentration of a reactant or product over time.
- **Rate Law:** An equation that relates the reaction rate to the concentrations of the reactants.
- **Activation Energy:** The minimum energy required for a reaction to occur.

Colligative Properties

- **Boiling Point Elevation:** The boiling point of a solution is higher than that of the pure solvent.
- **Freezing Point Depression:** The freezing point of a solution is lower than that of the pure solvent.
- **Osmotic Pressure:** The pressure required to stop the flow of solvent across a semipermeable membrane.

Atomic Structure

- **Atomic Number (Z):** The number of protons in the nucleus.
- **Mass Number (A):** The sum of protons and neutrons in the nucleus.
- **Isotopes:** Atoms of the same element with different numbers of neutrons.

Quantum Mechanics

- **Wave-Particle Duality:** Light and matter exhibit both wave-like and particle-like properties.
- **Schrodinger Equation:** A mathematical equation that describes the wave-like behavior of particles.
- **Atomic Orbitals:** Regions of space where there is a high probability of finding an electron.

Chemical Bonding

- **Covalent Bonding:** The sharing of electron pairs between atoms.
- **Electronegativity Difference:** A measure of the difference in electronegativity between two atoms.
- **Polarity:** The distribution of electrical charge across a molecule.

Chemical Equilibrium

- **Reaction Quotient (Q):** A ratio of product concentrations to reactant concentrations, used to determine the direction of a reaction.
- **Equilibrium Constant (K):** A ratio of product concentrations to reactant concentrations, used to determine the position of equilibrium.
- **Le Chatelier's Principle:** A system at equilibrium will shift to counteract any stress applied to it.

Chemical Kinetics

- **Rate Law:** An equation that relates the reaction rate to the concentrations of the reactants.
- **Activation Energy:** The minimum energy required for a reaction to occur.
- **Reaction Mechanism:** A series of steps that describe the overall reaction.

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course description outline and sample AP test questions It covers only the information tested on the new exam so you can make the most of your valuable study time Our targeted review focuses on the Big Ideas that will be covered on the exam Explanations of the AP Chemistry Labs are also included Expert Test taking Strategies This Crash Course presents detailed question level strategies for answering both the multiple choice and essay questions By following this advice you can boost your score in every section of the test Take REA's Online Practice Exam After studying the material in the Crash Course go to the online REA Study Center and test what you've learned Our practice exam features timed testing detailed explanations of answers and automatic scoring analysis The exam is balanced to include every topic and type of question found on the actual AP exam so you know you're studying the smart way Whether you're cramming for the test at the last minute looking for extra review or want to study on your own in preparation for the exams this is the study guide every AP Chemistry student must have When it's crucial crunch time and your Advanced Placement exam is just around the corner you need REA's Crash Course for AP Chemistry

Chemistry: 1,001 Practice Problems For Dummies (+ Free Online Practice) Heather Hattori, Richard H. Langley, 2014-03-11 Practice makes perfect and helps deepen your understanding of chemistry Every high school requires a course in chemistry and many universities require the course for majors in medicine engineering biology and various other sciences 1001 Chemistry Practice Problems For Dummies provides students of this popular course the chance to practice what they learn in class deepening their understanding of the material and allowing for supplemental explanation of difficult topics 1001 Chemistry Practice Problems For Dummies takes you beyond the instruction and guidance offered in Chemistry For Dummies giving you 1 001 opportunities to practice solving problems from the major topics in chemistry Plus an online component provides you with a collection of chemistry problems presented in multiple choice format to further help you test your skills as you go Gives you a chance to practice and reinforce the skills you learn in chemistry class Helps you refine your understanding of chemistry Practice problems with answer explanations that detail every step of every problem Whether you're studying chemistry at the high school college or graduate level the practice problems in 1001 Chemistry Practice Problems For Dummies range in areas of difficulty and style providing you with the practice help you need to score high at exam time

CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

5 Steps to a 5 AP Chemistry, 2010-2011 Edition John Moore, Richard H. Langley, 2010-01-01 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That's why we've created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You'll sharpen your subject knowledge strengthen your thinking skills and build your test taking

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The Teaching of Chemistry and Physics in the Secondary School Alexander Smith, Edwin Herbert Hall, 1902

Learning in the Fast Lane Chester E. Finn, Jr., Andrew E. Scanlan, 2021-05-11 More than three million high school students take five million Advanced Placement exams each May yet remarkably little is known about how this sixty year old privately run program has become one of U S education s greatest successes From its mid century origin as a tiny option for privileged kids from posh schools AP has also emerged as a booster rocket into college for hundreds of thousands of disadvantaged youngsters It challenges smart kids affects school ratings affords rewarding classroom challenges to great teachers tunes up entire schools and draws vast support from philanthropists education reformers and policymakers AP stands as America s foremost source of college level academics for high school pupils Praised for its rigor and integrity more than 22 000 schools now offer some or many of its thirty eight subjects from Latin to calculus art to computer science But challenges abound today as AP faces stiffening competition especially dual credit curriculum wars charges of elitism misgivings by elite schools and universities and the arduous work of infusing rigor into schools that lack it and academic success into young people unaccustomed to it In today s polarized climate can Advanced Placement maintain its lofty standards and overcome the hostility politics and despair that have sunk so many other bold education ventures Advanced Placement The Unsung Success Story of American Education is a unique account richly documented and thoroughly readable of the AP program in all its strengths and travails written by two of America s most respected education analysts

AP Government and Politics Research and Education Associa, Research & Education Association, Jack Hamilton, Robert F Gorman, Robert F. Gorman, Scott J Hammond, Scott J. Hammond, P Babbitt, Elliot Kalner, Wesley Phelan, J Hamilton, Gerald G. Watson, Keith Mitchell, 1992-12-09 A NEWER EDITION OF THIS TITLE IS AVAILABLE SEE ISBN 978 0 7386 0267 7 Get the AP college credits you ve worked so hard for Our savvy test experts show you the way to master the test and score higher This new and fully expanded edition examines all AP US a comparative review of the governments of England France the former Soviet Union and the People s Republic of China US political institutions public opinion voting behavior and more Features 3 full length practice exams with all answers thoroughly explained Follow up your study with REA s test taking strategies powerhouse drills and study

schedule that get you ready for test day DETAILS Comprehensive up to date subject review of every US these are referred to as the core countries tested on the exam For certain questions basic knowledge of the United States will be assumed II Free Response 50% of your grade This 100 minute section consists of four mandatory questions each of which accounts for one fourth of your total free response score You should allot roughly 25 minutes or one quarter of the total time in the free response segment for each essay Comparative Free Response questions may require you to compare one or two of the core countries Great Britain France China and the former Soviet Union with the developing nations of either India Mexico or Nigeria To do this you must be able to demonstrate knowledge of the politics of one of these developing nations Here s a breakdown of coverage on the Comparative exam Topics % of Exam I The Sources of Public Authority and Political Power 5 15% II Society and Politics 5 15% III The Relationship Between Citizen and State 5 15% IV Political and Institutional Frameworks 35 45% V Political Change 15 25% VI The Comparative Method 5 10% About the Review Sections As mentioned earlier this book includes two reviews one for United States Government and Politics the other for Comparative Government and Politics The United States Government and Politics Review covers all of the key information you ll need to score well on the United States exam These topics include Constitutional Framework The Federal Government Political Institutions and Special Interests Public Opinion and Voter Behavior Civil Rights and the Supreme Court We also provide a glossary for the United States Government and Politics exam Included are the key historical figures court cases programs laws etc that often appear on this AP exam The Comparative Review provides a thorough discussion of the material most often tested on the Comparative exam Special emphasis is placed on the governments and politics of Britain France The former Soviet Union The People s Republic of China A glossary for the Comparative Government and Politics exam enables you to brush up on terms that you are likely to encounter on this test Scoring the Exam After the AP administrations more than 1 700 college professors and secondary school teachers are brought together to grade the exams during the first two weeks of June These readers are chosen from around the United States for their familiarity with the AP program The Multiple Choice sections of the Comparative Government if the number is less than zero enter zero Scoring the Free Response Section The following guide explains typical free response scoring criteria Score Explanation of Score 8 9 The thesis is extremely well developed and is supported with concrete evidence all aspects of the question have been addressed thoroughly discussions presented are balanced 6 7 The thesis is defined and supported the evidence provided is very organized the essay may be slightly imbalanced with one strong argument and one weak argument and or discuss one topic more thoroughly than the next sporadic factual errors may appear 5 A basic argument or thesis is provided evidence given supports the argument or thesis but does not clearly connect with the argument or thesis only the formal facets of the question are dealt with and informal facets are not adequately covered not all aspects of the question are discussed 4 The thesis is not organized and is not referred to in the essay the essay is little more than a recounting of facts and events the essay may be overloaded with data

only one facet of the questions may be discussed numerous factual errors may appear 3 The thesis is weak evidence provided in support does not apply to the thesis factual errors are apparent 2 The thesis is very weak little or no factual evidence is provided to support the thesis irrelevant and inaccurate information appears 1 An attempt is made to answer the question but the support given is insignificant and the coverage of topics is incomplete 0 The question is not answered with any significance

Free Response Part II Comparative only Score Explanation of Score

5 The thesis is extremely well developed and is supported with concrete evidence all aspects of the question have been addressed thoroughly discussions are presented in a balanced way

4 The thesis is defined and supported the evidence provided is very organized the essay may be slightly imbalanced with one strong argument and one weak argument likewise one topic may be more thoroughly explored than another may be marred by sporadic factual errors

3 A basic argument or thesis is presented evidence given supports the argument or thesis but does not clearly connect with the argument or thesis only the formal facets of the question are dealt with and informal facets are not adequately covered not all aspects of the question are discussed

2 The thesis is weak evidence provided in support does not apply to the thesis factual errors are apparent

1 An attempt is made to answer the question but the support given is insignificant and the coverage of topics is incomplete

0 The question is not answered with any significance

It would be extremely helpful to find someone who is willing to score your essay your teachers or anyone who is familiar with the test material If you do ask the person to assign each of your U S and Comparative Part I essays a score of 0 to 9 For your Comparative Part II essays use the 0 to 5 scale If you must grade your own essays try to be objective In addition you may want to give your essays three different grades For instance if you feel you did well try giving the essay a score of 5 6 or 7 to represent the various scores you may receive By underestimating what your score may be you are more likely to receive a better score on the actual exam Use the following formulae to determine your raw score for the Free Response section

United States Exam Free Response Response 1 score x 1 66 raw score Response 2 score x 1 66 raw score Response 3 score x 1 66 raw score Response 4 score x 1 66 raw score

Comparative Exam Response 1 score x 1 66 raw score Response 2 score x 1 66 raw score Response 3 score x 3 raw score Response 4 score x 3 raw score

The Composite Score

Once you have obtained your raw scores for both the Multiple Choice and the Free Response sections add the scores together to get your composite score

United States Exam Multiple Choice raw score Free Response raw score composite score round to nearest whole number

Score Essay 1 Score Essay 2 Score Essay 3 Score Essay 4 raw score

Comparative Exam Multiple Choice raw score Free Response raw score composite score round to nearest whole number

Now compare your composite score with the scale below

Composite Score	AP Grade
88-120	5
74-87	4
54-73	3
35-53	2
0-34	1

AP grades are interpreted as follows

- 5 extremely well qualified
- 4 well qualified
- 3 qualified
- 2 possibly qualified
- 1 no recommendation

Scores that Earn College Credit and or Advanced Placement Most colleges grant students who earn a 3 or above college credit and or advanced placement You should check with your school guidance office about specific college requirements

Studying for Your AP Examination It is never too early to start studying The earlier you begin the more time you will have to sharpen your skills Do not procrastinate Cramming is not an effective way to study since it does not allow you the time needed to learn the test material It is very important for you to choose the time and place for studying that works best for you Some students may set aside a certain number of hours every morning to study while others may choose to study at night before going to sleep Other students may study during the day while waiting on a line or even while eating lunch Only you can determine when and where your study time will be most effective But be consistent and use your time wisely Work out a study routine and stick to it When you take the practice exam s try to make your testing conditions as much like the actual test as possible Turn your television and radio off and sit down at a quiet table free from distraction Make sure to time yourself As you complete the practice test s score your test s and thoroughly review the explanations to the questions you answered incorrectly but do not review too much during any one sitting Concentrate on one problem area at a time by reviewing the question and explanation and by studying our review s until you are confident that you completely understand the material Since you will be allowed to write in your test booklet during the actual exam you may want to write in the margins and spaces of this book when practicing However do not make miscellaneous notes on your answer sheet Mark your answers clearly and make sure the answer you have chosen corresponds to the question you are answering Keep track of your scores By doing so you will be able to gauge your progress and discover general weaknesses in particular sections You should carefully study the reviews that cover the topics causing you difficulty as this will build your skills in those areas To get the most out of your studying time we recommend that you follow the Study Schedule which corresponds to the exam you are taking It details how you can best budget your time If you are taking both exams do not try to study for each at the same time Try alternating days by studying for the United States exam one day and the Comparative exam the next

Test Taking Tips Although you may be unfamiliar with tests such as the Advanced Placement exams there are many ways to acquaint yourself with this type of examination and help alleviate your test taking anxieties Listed below are ways to help yourself become accustomed to the AP exam some of which may also be applied to other standardized tests Become comfortable with the format of the AP Examination in Government and Politics that you are taking When you are practicing to take the exam s simulate the conditions under which you will be taking the actual test s You should practice under the same time constraints as well Stay calm and pace yourself After simulating the test only a couple of times you will boost your chances of doing well and you will be able to sit down for the actual test much more confidently Know the directions and format for each section of the exam Familiarizing yourself with the directions and format of the different test sections will not only save you time but will also ensure that you are familiar enough with the AP exam to avoid nervousness and the mistakes caused by being nervous Work on the easier questions first If you find yourself working too long on one question make a mark next to it in your test booklet and continue After you have answered all of the questions that you can go back to the ones you have

skipped Use the process of elimination when you are unsure of an answer If you can eliminate three of the answer choices you have given yourself a fifty fifty chance of getting the item correct since there will only be two choices left from which to make a guess If you cannot eliminate at least three of the answer choices you may choose not to guess as you will be penalized one quarter of a point for every incorrect answer Questions not answered will not be counted Be sure that you are marking your answer in the circle that corresponds to the number of the question in the test booklet Since the multiple choice section is graded by machine marking the wrong answer will throw off your score

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