
ACS Gen Chem 1 Exam

Ion - A species of an element in which the number of electrons does not equal the number of protons.

Isotope - Each isotope of an atom has a different number of neutrons.

Isoelectric - Two atoms with the same charge.

Isotopic - Two atoms with the same number of neutrons.

Transition Metal Orbitals - The 4s orbital is filled before the 3d orbital, but listed after.

Paramagnetic - Atoms that possess a permanent magnetic charge due to the presence of unpaired electrons.

Electropositive - Tending to lose electrons and form positive ions.

Tetrahedral Bond Angle - 109.5°

Resonance Structures - Two equally correct arrangements of one molecule's electrons; generally the molecule exists as an average of the two.

Polar Molecule - A molecule that is mostly positive on one side and mostly negative on the other.

sp^3 Orbital - An orbital composed of one s orbital and three p orbitals; can hold eight total electrons. Tetrahedral.

Trigonal Planar - The way that three objects will arrange themselves around a central atom. Non-polar.

Free Radicals - Compounds with an odd number of electrons in their Lewis Structure. Eg. NO_2

Incomplete Octets - Occur in elements that can form stable bonds with less than eight electrons. Eg. BF_3

Expanded Octets - Occur in the third row of the periodic table and beyond in elements that can form stable bonds with more than eight electrons. Eg. XeF_2

Hydrogen Bonding - When hydrogen is attracted to electronegative atoms.

Acs Review Guide

M Woodhall



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Stereoisomerism Nucleophilic Substitutions and Eliminations Electrophilic Additions Nucleophilic Addition at Carbonyl Groups Nucleophilic Substitution at Carbonyl Groups Enols and Enolate Ion Reactions Electrophilic and Nucleophilic Aromatic Substitution Free Radical Substitutions and Additions Oxidations and Reductions Spectroscopy Synthesis and Analysis Practice Questions Practice makes perfect Detailed Answer Explanations Figure out where you went wrong and how to improve Studying can be hard We get it That s why we created this guide with these great features and benefits Comprehensive Review Each section of the test has a comprehensive review created by Test Prep Books that goes into detail to cover all of the content likely to appear on the test ACS Organic Chemistry Practice Test Questions We want to give you the best practice you can find That s why the Test Prep Books practice questions are as close as you can get to the actual test Answer Explanations Every single problem is followed by an answer explanation We know it s frustrating to miss a question and not understand why The answer explanations will help you learn from your mistakes That way you can avoid missing it again in the future Test Taking Strategies A test taker has to understand the material that is being covered and be familiar with the latest test taking strategies These strategies are necessary to properly use the time provided They also help test takers complete the test without making any errors Test Prep Books has provided the top test taking tips Customer Service We love taking care of our test takers We make sure that you interact with a real human being when you email your comments or concerns Anyone planning to take this exam should take advantage of this Test Prep Books study guide Purchase it today to receive access to ACS Organic Chemistry review materials ACS Organic Chemistry practice test questions Test taking strategies

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The ACS Style Guide Janet S. Dodd, 1997 Guidelines from ACS to help authors and editors in preparing scientific texts

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