

Acid-Base Equilibrium Homework Part II

Part A

1. Calculate the pH and pOH of a solution containing 7.30 grams of HCl in 1.0 L of solution.

$$7.30 \text{ g} \times \frac{1 \text{ mol}}{36.46 \text{ g}} = 0.200 \text{ mol} / 1.0 \text{ L} = 0.20 \text{ M HCl}$$

Because HCl is a strong acid, it completely dissociates, so $[\text{HCl}] = [\text{H}_3\text{O}^+] = 0.20 \text{ M}$

$$\text{pH} = -\log[0.20] = 0.70$$

$$\text{pOH} = 14 - \text{pH} = 13.30$$

2. Calculate the pH of a 0.10 M nitric acid (HNO_3) solution.

Because HNO_3 is a strong acid, the acid completely dissociates and the concentration of the acid is equal to the concentration of H^+ at equilibrium.

$$[\text{H}_3\text{O}^+] = 0.10 \text{ M}$$

$$\text{pH} = -\log(0.10) = 1$$

3. Calculate the pH of a $5.0 \times 10^{-3} \text{ M}$ solution of sodium hydroxide.

Sodium hydroxide is a strong base, which means it completely dissociates, so $[\text{NaOH}] = [\text{OH}^-]$.

$$[\text{OH}^-] = 5.0 \times 10^{-3}$$

$$\text{pOH} = -\log(5.0 \times 10^{-3}) = 1.3$$

$$\text{pH} = 14 - 1.3 = 12.7 = 13 \text{ (2 sig figs)}$$

Part B

4. Decide whether the statements are true about strong acids, weak acids, both, or neither.

- a. Always dissociates leaving behind at least some hydronium ion **both**
- b. The initial concentration of the acid is equivalent to the final concentration of hydronium ions **strong only**
- c. The pH is easily calculated **strong only**
- d. The pH must be calculated by the ICE method **weak only**
- e. The acid completely dissociates in water **strong only**
- f. The equilibrium concentration of undissociated acid is approximately 0 **strong only**
- g. K_a is very small **weak only**

5. Write the K_a expression for the following acids:

- a. $\text{HC}_2\text{H}_3\text{O}_2$



$$K_a = \frac{[\text{C}_2\text{H}_3\text{O}_2^-][\text{H}_3\text{O}^+]}{[\text{HC}_2\text{H}_3\text{O}_2]}$$

- b. H_2CO_3



$$K_a = \frac{[\text{HCO}_3^-][\text{H}_3\text{O}^+]}{[\text{H}_2\text{CO}_3]}$$

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 - The Rise of Digital Reading Acid Base Homework Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Homework Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Homework Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Homework Answers
 - Personalized Recommendations
 - Acid Base Homework Answers User Reviews and Ratings
 - Acid Base Homework Answers and Bestseller Lists
5. Accessing Acid Base Homework Answers Free and Paid eBooks
 - Acid Base Homework Answers Public Domain eBooks
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 - Acid Base Homework Answers Budget-Friendly Options

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 - Acid Base Homework Answers Compatibility with Devices
 - Acid Base Homework Answers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Acid Base Homework Answers
 - Highlighting and Note-Taking Acid Base Homework Answers
 - Interactive Elements Acid Base Homework Answers
8. Staying Engaged with Acid Base Homework Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Homework Answers
9. Balancing eBooks and Physical Books Acid Base Homework Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Homework Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acid Base Homework Answers
 - Setting Reading Goals Acid Base Homework Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid Base Homework Answers
 - Fact-Checking eBook Content of Acid Base Homework Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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