

Chemistry 162 Worksheet 3

Name _____ ID _____

$$S = k_m(P), [A]_t = -kt + [A]_{in}, P_{total} = X_A(P_A) + X_B(P_B), 1/[A]_t = kt + 1/[A]_{in}, \Delta T = K_{cal}(m)$$

$$q = m(C_p)\Delta T, \ln[A]_t = -kt + \ln[A]_{in}, P_{mixture} = MRT, R = 0.08206 \text{ atm}\cdot\text{L/mol}\cdot\text{K}$$

1.) Given the following balanced equation, determine the rate of reaction with respect to $[\text{H}_2]$.

$$\text{For a Reaction: } a \text{A} + b \text{B} \rightarrow c \text{C} + d \text{D} \quad \text{Rate} = -\frac{1}{a} \frac{\Delta[\text{A}]}{\Delta t} = -\frac{1}{b} \frac{\Delta[\text{B}]}{\Delta t} = +\frac{1}{c} \frac{\Delta[\text{C}]}{\Delta t} = +\frac{1}{d} \frac{\Delta[\text{D}]}{\Delta t}$$

$$\text{A) Rate} = +\frac{3 \Delta[\text{H}_2]}{\Delta t} \quad \text{B) Rate} = -\frac{3 \Delta[\text{H}_2]}{\Delta t} \quad \text{C) Rate} = +\frac{1}{3} \frac{\Delta[\text{H}_2]}{\Delta t} \quad \text{D) Rate} = -\frac{1}{3} \frac{\Delta[\text{H}_2]}{\Delta t}$$

2.) Write a balanced reaction for which the following rate relationships are true.

$$\text{Rate} = -\frac{1}{2} \frac{\Delta[\text{N}_2\text{O}_5]}{\Delta t} = \frac{1}{4} \frac{\Delta[\text{NO}_2]}{\Delta t} = \frac{\Delta[\text{O}_2]}{\Delta t}$$



3.) Give the characteristic of a second order reaction having only one reactant.

A) The rate of the reaction is not proportional to the concentration of the reactant.

B) The rate of the reaction is proportional to the square of the concentration of the reactant.

C) The rate of the reaction is proportional to the square root of the concentration of the reactant.

D) The rate of the reaction is proportional to the natural logarithm of the concentration of the reactant.

4.) What are the units of k in the following rate law?

$$\text{Rate} = k[\text{X}][\text{Y}]^2$$



$$k_{(2^{\text{nd}} \text{ order})} = \frac{1}{\text{M}^{2-1} \cdot \text{time}} = \frac{\text{L}^{\text{mol}-1}}{\text{mol}^{\text{L}-1} \cdot \text{time}}$$

$$\text{A) } \frac{1}{\text{M s}^2} \quad \text{B) } \frac{1}{\text{M}^2 \text{ s}} \quad \text{C) } \text{M}^2 \text{ s} \quad \text{D) } \frac{\text{M}^2}{\text{s}} \quad \text{E) } \frac{1}{\text{M}^3 \text{ s}}$$

5.) Determine the rate law and the value of k for the following reaction using the data provided.

$\text{CO}(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow \text{COCl}_2(\text{g})$	$[\text{CO}]_i \text{ (M)}$	$[\text{Cl}_2]_i \text{ (M)}$	Initial Rate ($\text{M}\cdot\text{s}^{-1}$)
	0.25	0.40	0.696
	0.25	0.80	1.97
	0.50	0.80	3.94

$$\text{A) Rate} = 11 \text{ M}^{-3/2}\text{s}^{-1} [\text{CO}][\text{Cl}_2]^{3/2}$$

$$\text{B) Rate} = 36 \text{ M}^{-1}\text{s}^{-1} [\text{CO}][\text{Cl}_2]^{2.8}$$

$$\text{C) Rate} = 17 \text{ M}^{-2}\text{s}^{-1} [\text{CO}][\text{Cl}_2]^2$$

$$\text{D) Rate} = 4.4 \text{ M}^{-1/2}\text{s}^{-1} [\text{CO}][\text{Cl}_2]^{1/2}$$



Use "Method of Initial Rates" to find the order of the reaction with respect to each reactant (there are multiple ways to solve this):

Comparing Exp. 2 and Exp. 3, when $[\text{CO}]$ doubles and $[\text{Cl}_2]$ is constant, the rate doubles. This means the reaction is 1st order with respect to $[\text{CO}]$. Comparing Exp. 1 and Exp. 2, it is harder to see the relationship

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