

Acid Base Equilibria

22 pages

Titrations

- It needs acid and the concentration of the acid is a base
- If the acid/base of **known concentration** is the **titrant**
- If the acid/base of **unknown concentration** the **titrant** is the **analyte** to be added
- If adding base to solution into the conical flask until the solution is **neutralised** - stops to turn from the red point
- If **indicator** changes are done then the **indicator** of the reactions are required to **neutralise** the acid with the strong base - **indicator** colour is **neutral** colour is **blue**
- If **titrant** until one gets **1 concentration** - **indicator** - **indicator** - **indicator**

Titration curve

When pH changes rapidly, added base is neutralised.

Strong acid - strong base



- If **strong acid** is in the conical flask and base is being added to it
- If **strong base** is being added to the solution in the conical flask, then the **titrant** is being added to it
- If **acid** is at pH 1 and **base** is at pH 14, then the **titrant** is being added to it
- If **base** is at pH 14 and **acid** is at pH 1, then the **titrant** is being added to it

Strong acid - weak base



Bronsted Lowry

Acid is a proton donor
Base is a proton acceptor

- If **strong acid** and **strong base** are added, then the **strong acid** and **strong base** are added to the solution



- If **acid** is the **acid** and **base** is the **base** in the **strong acid** and **base** is the **base**
- If **acid** and **base** are **conjugate acid** and **base**
- If **acid** and **base** are **conjugate acid** and **base**

- If **strong acid** and **strong base** are added to the solution
- If **strong base** and **strong acid** are added to the solution



- If **strong acid** and **strong base** are added to the solution
- If **strong acid** and **strong base** are added to the solution
- If **strong acid** and **strong base** are added to the solution

Conjugate pairs

Identify the transfer of a proton



- If **acid** is the **acid** and **base** is the **base** in the **strong acid** and **base** is the **base**
- If **acid** and **base** are **conjugate acid** and **base**
- If **acid** and **base** are **conjugate acid** and **base**

Acid-base reactions

ionic product of water

Water is in equilibrium with its ions
Water dissociates into **hydroxide** and **hydrogen** ions. $[OH^-] = [H^+] = 10^{-7}$

Water is in equilibrium with its ions

Water is in equilibrium with its ions

- If **strong acid** and **strong base** are added to the solution
- If **strong acid** and **strong base** are added to the solution
- If **strong acid** and **strong base** are added to the solution
- If **strong acid** and **strong base** are added to the solution

$[H^+] = 10^{-7}$ and $[OH^-] = 10^{-7}$ in pure water



Example

What is the pH of a 0.1M solution of HCl?

If **strong acid** and **strong base** are added to the solution



Water is in equilibrium with its ions

Example

- If **strong acid** and **strong base** are added to the solution
- If **strong acid** and **strong base** are added to the solution
- If **strong acid** and **strong base** are added to the solution



Acid Base Equilibria Study Guide Answers

Neil E. Schore



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