



Some Basic Concepts of Chemistry



$$\text{Velocity} = \frac{\text{Distance}}{\text{Time}}$$



$$\text{Acceleration} = \frac{\text{Velocity}}{\text{Time}}$$

$$1 \text{ amu} = 1.66056 \times 10^{-24} \text{ gram}$$

Atomic mass unit



$$\text{Density} = \frac{\text{Mass}}{\text{Volume}}$$



$$\text{Pressure} = \frac{\text{Force}}{\text{Area}}$$



$$^{\circ}\text{F} = \frac{9}{5} (^{\circ}\text{C}) + 32$$

$$\text{K} = ^{\circ}\text{C} + 273.15$$

$^{\circ}\text{F}$ - Fahrenheit
 $^{\circ}\text{C}$ - Celsius
 K - Kelvin



$$\text{Number of grams of atoms} = \frac{\text{Weight (in grams)}}{\text{gram atomic mass}}$$



$$\text{Molecular Mass } (A_x B C_y) = x \times [\text{Atomic mass of A}] + [\text{Atomic mass of B}] + y \times [\text{Atomic mass of C}]$$

OR



$$\text{Molecular Mass} = \frac{\text{Actual mass of a molecule of matter}}{\frac{1}{12} \times \text{mass of one } ^{12}\text{C atom}}$$



$$\text{Gram Molecular No.} = \frac{\text{Mass (in grams)}}{\text{Molecular weight}}$$



$$\text{Equivalent weight of the element} = \frac{\text{Atomic weight of the element}}{\text{valency}}$$



$$\text{Equivalent weight of an element} = \frac{\text{formula weight of salt}}{\text{valency of the element in the salt}}$$



$$\text{Equivalent weight of an acid} = \frac{\text{Molecular weight of acid}}{\text{number of replaceable H}^+ \text{ ions}}$$



$$\text{Equivalent weight of a base} = \frac{\text{Molecular weight of base}}{\text{number of replaceable OH}^- \text{ ions}}$$



$$\text{Equivalent weight of a salt} = \frac{\text{Molecular weight of salt}}{\text{total positive or negative charge}}$$



$$\text{Equivalent weight of an oxidizing agent} = \frac{\text{Molecular weight of oxidizing agent}}{\text{change in oxidation number of the element in the molecule}}$$



$$\text{Equivalent weight of a reducing agent} = \frac{\text{Molecular weight of reducing agent}}{\text{change in oxidation number of the element in the molecule}}$$

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Kenneth Tobin,Ashraf Shady



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