

Acids and Bases Worksheet 1

1. Finish the following reactions when acids are added to water.



2. What are the conjugate bases of these acids?

original acid	conjugate base
HNO_3	
H_2O	
H_3O^+	
H_2SO_4	
HBr	
HCO_3^-	

3. What are the conjugate acids of these bases?

original base	conjugate acid
OH^-	
H_2O	
HCO_3^-	
SO_4^{2-}	
ClO_4^-	

4. Which of the following represent conjugate acid-base pairs?



5. Calculate the $[\text{H}^+]$ in a solution in which $[\text{OH}^-] = 2.0 \times 10^{-2} \text{ M}$. Is this solution acidic, neutral, or basic?

6. Calculate the $[\text{OH}^-]$ in a solution in which $[\text{H}^+] = 3.99 \times 10^{-5} \text{ M}$. Is this solution acidic, neutral, or basic?

Acids And Bases Study Guide Answer Key

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Bell, Terry Thompson, Peggie L.
Williamson**

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