

Pre-Laboratory Assignment: Acid-Base Titrations

Name _____ Section _____ Date _____

1. Find in your textbook the dilution formula and calculate the amount of 6M NaOH required to prepare 500 ml of 0.25 M NaOH. 2 decimals

$$M_1 = 6M \text{ NaOH}$$

$$V_1 = ?$$

$$M_2 = 0.25M$$

$$V_2 = 500mL$$

$$V_1 = \frac{M_2 V_2}{M_1}$$

$$V_1 = \frac{0.25 \times 500}{6.0} = 20.83mL$$

2. Repeat the calculation for the amount of 6 M HCl required to prepare 300 mL of 0.35 M HCl. 2 decimals

$$M_1 = 6M \text{ HCl}$$

$$V_1 = ?$$

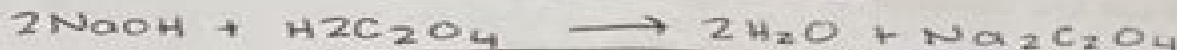
$$V_2 = 300mL$$

$$M_2 = 0.35M$$

$$V_1 = \frac{300 \times 0.35}{6.0} = 17.50$$

3. Calculate the mass of oxalic acid dihydrate ($H_2C_2O_4 \cdot 2H_2O$) required to react with 15. mL of 0.25 M sodium hydroxide solution. (This is the calculation, needed in Part III, of the sample size of your primary standard for the standardization of NaOH.) 1 decimal

$$0.015 \times 0.25 = 0.00375 \text{ mol NaOH}$$



$$0.00375 \text{ mol NaOH} \times \frac{1}{2} = 0.001875 \text{ mol } H_2C_2O_4$$

$$\text{mol } H_2C_2O_4 \cdot 2H_2O = 0.001875$$

$$0.001875 \times 126.07 = 0.236g$$

$$\text{Mass } H_2C_2O_4 \cdot 2H_2O = 0.2g$$

Acid Base Titration Pre Lab Answers Uic

Kendon Douglas Smith

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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1. Understanding the eBook Acid Base Titration Pre Lab Answers Uic
 - The Rise of Digital Reading Acid Base Titration Pre Lab Answers Uic
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titration Pre Lab Answers Uic
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Titration Pre Lab Answers Uic
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titration Pre Lab Answers Uic
 - Personalized Recommendations
 - Acid Base Titration Pre Lab Answers Uic User Reviews and Ratings

- Acid Base Titration Pre Lab Answers Uic and Bestseller Lists
- 5. Accessing Acid Base Titration Pre Lab Answers Uic Free and Paid eBooks
 - Acid Base Titration Pre Lab Answers Uic Public Domain eBooks
 - Acid Base Titration Pre Lab Answers Uic eBook Subscription Services
 - Acid Base Titration Pre Lab Answers Uic Budget-Friendly Options
- 6. Navigating Acid Base Titration Pre Lab Answers Uic eBook Formats
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 - Acid Base Titration Pre Lab Answers Uic Enhanced eBook Features
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 - Highlighting and Note-Taking Acid Base Titration Pre Lab Answers Uic
 - Interactive Elements Acid Base Titration Pre Lab Answers Uic
- 8. Staying Engaged with Acid Base Titration Pre Lab Answers Uic
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Titration Pre Lab Answers Uic
- 9. Balancing eBooks and Physical Books Acid Base Titration Pre Lab Answers Uic
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Titration Pre Lab Answers Uic
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid Base Titration Pre Lab Answers Uic
 - Setting Reading Goals Acid Base Titration Pre Lab Answers Uic
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid Base Titration Pre Lab Answers Uic
 - Fact-Checking eBook Content of Acid Base Titration Pre Lab Answers Uic
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
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