

BrainPOP- pH Scale

Name _____

Visit the main page of the **Kid Zone** at <http://sciencespot.net/> to find the link for BrainPOP.

1. The term _____ refers to the chemical potential of hydrogen.
2. The pH _____ measures how _____ or basic a substance on a scale of 1 to 14. _____ measure from 1 to 7, while the _____ or alkaline side measures from 7 to 14.

3. Identify each as an acid (A) or a base (B).

____ Lemon juice ____ Soap ____ Battery acid ____ Hydrochloric acid
____ Toothpaste ____ Vinegar ____ Baking soda ____ Floor cleaner

4. Really strong acids and bases are found at the _____ of the scale, while those near the _____ are weaker. Right in the middle at exactly 7, or pH _____, you have pure water.

5. Acids contain _____ ions (H^+), while bases contain _____ (OH^-) ions. Ions are atoms that have either lost or gained _____. In bases you find hydroxyl ions, which have _____ electrons. Hydrogen ions found in acids are _____ that have had an electron knocked off, which is why these ions are always trying to get the lost electrons back. When you put a metal in acid, the metal starts to _____.

6. Bases have the power to _____ acids. The hydroxyl ion _____ its electron with the hydrogen ion. The H^+ from the hydrogen ion and the OH^- from the hydroxyl ion bond together to form H_2O , which is _____ with pH neutral. Different types of _____ are formed depending on the types of acids and bases involved in the reaction.

7. When Moby poured the baking soda (a base) into the vinegar (an acid), a reaction occurred. Salt, water, and _____ were produced by the reaction.

8. _____ are substances that can block changes in their pH for period of time. You can find them on buffered aspirin, which keeps your _____ from dissolving the medicine too quickly. Buffers are made by combining _____ acids and bases. Buffers also are found in our body to keep our the pH of our _____ at or around 7.4.

9. Most science labs have pH paper, which is paper that has been soaked in a special chemical _____. They turn _____ in an acid and _____ in a base. By comparing the color of the pH paper to a chart, you can determine how _____ an acid or base is.

10. Constant reactions between acids and bases keep our _____ from being too acidic or too alkaline.

Word Bank:

Acidic
Acids
Atoms
Bases
Blood
Blue
Buffers
Carbon dioxide
Center
Dissolve
Electrons
Ends
Environment
Gained
Hydrogen
Hydroxyl
Indicator
Neutral
Neutralize
pH
Red
Salts
Scale
Shares
Stomach acid
Strong
Water
Weak

Asnwer Key For Brainpop Ph Scale

NARAYAN CHANGDER



Answer Key For Brainpop Ph Scale:

The pH Scale Mary Griffin, 2018-12-15 The pH scale measures how acidic or basic a substance is ranging from 0 to 14 Readers will learn how certain substances rank on the pH scale what happens when acids and bases are mixed and how water can make a substance either acidic or basic These significant science concepts are discussed in approachable text and supported by motivating fact boxes charts images and photographs

Ph Measurement. Specification for Ph Scale British Standards Institute Staff, British Standards Institution, 1984-12-31 pH measurement pH Range statistics Accuracy Grades quality Standard solutions Phthalates Potassium organic compounds *PH Scale*, 2006

PH Scale British Standards Institution, 1961

The PH Scale Marie Roesser, 2025-07-30 The pH scale measures how acidic or basic a substance is ranging from 0 to 14 This chemistry concept has real world applications For example strawberries grow best in soil with a specific pH Safe drinking water must have a pH within a certain range too Readers will learn more about pH the scale how substances are tested and how pH can change in this accessible volume a valuable addition to any elementary science collection Essential curricular information is supported by Learn More fact boxes labeled diagrams and a summarizing graphic organizer

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Ph Scale And Acid-base Titrations/standardization Postma Roberts Hollenberg, James M. Postma, Julian Roberts, J. Leland Hollenberg, 2000-01-15

Specification for PH Scale British Standards Institution, 1950

The Meaning and Calibration of the PH Scale, 1933

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