

## Chemistry Unit 3 Test Review

### 5.1 Mole and Avogadro's Constant

A **particle** is very small and can be:

- A single atom (Ca)
- A molecule with 2 or more atoms (molecular compound – H<sub>2</sub>O)
- An ion with a positive or negative charge (CO<sub>3</sub><sup>2-</sup>)
- A formula unit with two or more ions (ionic compound – NaCl)

#### The Mole

- Chemists use the mole to represent a specific number of particles
- The number of particles it represents is called the Avogadro constant ( $6.02 \times 10^{23}$  particles/mol)
- The unit for mole is mol

#### Definitions

1. Mole: One mole of any entity is  $6.02 \times 10^{23}$  of the particles (atom, molecule, ion formula unit); the symbol is n; the unit is mol
2. Avogadro's constant: The number of particles in one mole,  $6.02 \times 10^{23}$ ; the symbol is N<sub>A</sub>; the unit is particles/mol
3. Particles:
  - Atom or element (Eg. Zn)
  - Molecule or covalent compound (Eg. H<sub>2</sub>O)
  - Ions or monatomic/polyatomic ions (Eg. Zn<sup>+2</sup> and CO<sub>3</sub><sup>2-</sup>)
  - Formula unit or ionic compound (Eg. LiCl)

### 5.2 Molar Mass

#### Basics: The Mass and The Mole

- The mass of carbon-12 is 12 atomic mass unit (12 u). Every other element has mass relative to one atom of carbon-12
- 1 mole of carbon-12 =  $6.02 \times 10^{23}$  atoms of C-12 = 12g
- The average atomic mass of any element, in atomic mass units, is numerically equivalent to the mass of one mole of atoms ( $6.02 \times 10^{23}$  atoms) of that element

**Molar Mass:** The mass of one mole of its particle – atoms, molecules, ions, or formula units, in grams per mole (g/mol)

### 6.1 Percentage Composition

#### Law of Definite Proportions:

- Elements in a chemical compound are always present in the same proportions by mass
- *Law of constant composition:* all molecules or formula units of a compound will have the same chemical composition

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