

Activity 10

Classifying reactions

Can you classify these?

Type of chemical reactions and their descriptions



Introduction: Chemical reactions can be **classified** and **described** using a set of words. An equation is a shorthand way of writing a chemical reaction. Each side of the equation shows the reactants and products. A compound is a substance made of more than one kind of atom. Each side of the equation shows what reacts.

Question: How are chemical reactions classified?

1. **Match:** Match chemical reactions with the classified as one of four types. Using the chemical equations in the boxes on the right, match the following definitions to the type of reaction.

- ☒ **A** One reaction involves two or more products.
- ☒ **A** One reaction involves two or more reactants.
- ☒ **A** Two or more reactants combine to form one product.
- ☒ **A** Two compounds react to form two different compounds.
- ☒ **A** A compound reacts with an element to form one compound and one different element.

- A** **Reduction**
- A** **Decomposition**
- C** **Single replacement**
- B** **Double replacement**
- A** **Combustion**

2. **Match:** Match each of the chemical equations below. Some reactions may already be classified in the space on the right. Match the reaction as a synthesis, decomposition, single replacement, or double replacement reaction.

- A** $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- B** $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- C** $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- D** $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- E** $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

- Double**
- Single**
- Double**
- Double**
- Double**

Balancing Equations Gizmo Key

Randy L. Bell, Joe Garofalo

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