

TESTS

EXPLORE CHEMISTRY WITH CHEMISTRY

DESCRIBING CHEMICAL REACTIONS

SOLUTIONS TO THE MODULE 7 REVIEW QUESTIONS

1. Elements cannot undergo decomposition reactions because elements have nothing to decompose into.
2. a. This equation represents a synthesis reaction because 2 elements come together to produce 1 compound.
b. This equation represents a formation reaction because 2 elements come together to produce 1 compound.
3. None of these
4. None of these
5. This equation represents a decomposition reaction because a single compound is breaking down into its elements.
6. The only equation that does not represent a formation reaction is (b). The equation for (b) is a formation equation because it starts with 2 reactants and makes only 1 in the end. Remember, reactions do not have to start with elements in order to be formation equations; we just assume they do when we write them ourselves.
7. The only equation that does not represent a combustion reaction is (b). The equation for (b) is a decomposition equation. Even though it produces CO_2 and H_2O , it does not have O_2 as a reactant, so it is not combustion.
8. Complete combustion produces carbon dioxide and water. Incomplete combustion produces carbon monoxide and water or carbon and water.

...ment units, and 4 pts for the answer. If the student, this is the same as gold, but we will go units divided by volume, it is known density and is known, that there are conflicting units. The units are given in kg. Thus, before we solve the units kg to g.

... can use the equation for density to solve for volume

... 25 g
... 100 g
... 100 g

... each answer, minus 2 pts for every wrong element found
... are atomic elements, so this has one nitrogen, 2s, and fluorine, 1s.

The equation is balanced.

... 25 g
... 100 g
... 100 g

2 pts for the correct decomposition, 4 pts for the equation, and 4 pts for the balancing of: $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$

... 4 pts for the correct element composition, 4 pts for the equation, and 4 pts for the balancing of: $\text{C}_2\text{H}_6 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$

... 4 pts for the correct element composition, 4 pts for the equation, and 4 pts for the balancing of: $\text{C}_2\text{H}_6 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$

4 pts for the correct equation is balanced, then the same number of atoms appear on both sides. They would mean the same in the same, whether we are talking about mass or products. That's another way of saying the law of mass conservation.

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Jay L. Wile



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Table of Contents Apologia Chemistry Solutions And Test Manual

1. Understanding the eBook Apologia Chemistry Solutions And Test Manual
 - The Rise of Digital Reading Apologia Chemistry Solutions And Test Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Apologia Chemistry Solutions And Test Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Apologia Chemistry Solutions And Test Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Apologia Chemistry Solutions And Test Manual
 - Personalized Recommendations
 - Apologia Chemistry Solutions And Test Manual User Reviews and Ratings

- Apologia Chemistry Solutions And Test Manual and Bestseller Lists
- 5. Accessing Apologia Chemistry Solutions And Test Manual Free and Paid eBooks
 - Apologia Chemistry Solutions And Test Manual Public Domain eBooks
 - Apologia Chemistry Solutions And Test Manual eBook Subscription Services
 - Apologia Chemistry Solutions And Test Manual Budget-Friendly Options
- 6. Navigating Apologia Chemistry Solutions And Test Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Apologia Chemistry Solutions And Test Manual Compatibility with Devices
 - Apologia Chemistry Solutions And Test Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Apologia Chemistry Solutions And Test Manual
 - Highlighting and Note-Taking Apologia Chemistry Solutions And Test Manual
 - Interactive Elements Apologia Chemistry Solutions And Test Manual
- 8. Staying Engaged with Apologia Chemistry Solutions And Test Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Apologia Chemistry Solutions And Test Manual
- 9. Balancing eBooks and Physical Books Apologia Chemistry Solutions And Test Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Apologia Chemistry Solutions And Test Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Apologia Chemistry Solutions And Test Manual
 - Setting Reading Goals Apologia Chemistry Solutions And Test Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Apologia Chemistry Solutions And Test Manual
 - Fact-Checking eBook Content of Apologia Chemistry Solutions And Test Manual
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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