

Задача 11.
Слабые и сред. кислоты

① Дано:

$$C_{\text{HNO}_2} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-4}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-3} = 6 - \lg 2,8 = 6 - 0,45 = 5,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HNO}_2}}$$

$$= \sqrt{7,9 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-3} \text{ моль/л}$$

② Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}^+} = C_{\text{OH}^-} = 2,5 \cdot 10^{-3}$$

$$K_{\text{Pr}} = [C_{\text{H}^+}][C_{\text{OH}^-}] = 2,5 \cdot 10^{-3} \cdot 2,5 \cdot 10^{-3} = 6,25 \cdot 10^{-6}$$

$$= 6,25 \cdot 10^{-6}$$

3 - неизвестная моль/л

$$S = \sqrt[3]{\frac{K_{\text{Pr}}}{4}} = \sqrt[3]{\frac{6,25 \cdot 10^{-6}}{4}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[C_{\text{H}^+}] = \frac{S}{2} = \frac{1,16 \cdot 10^{-2}}{2} = 5,8 \cdot 10^{-3} \text{ моль/л}$$

$$[C_{\text{H}^+}] = \frac{10^{-14}}{[C_{\text{OH}^-}]} = \frac{10^{-14}}{5,8 \cdot 10^{-3}} = 1,72 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 1,72 \cdot 10^{-12} = 12 - \lg 1,72 = 11,76$$

③ Дано:

$$V_{\text{HNO}_2} = V_{\text{NaOH}}$$

$$C_{\text{HNO}_2} = 0,1 \text{ м}$$

$$V_{\text{NaOH}} = 10 \text{ мл}$$

$$C_{\text{NaOH}} = 0,1 \text{ м}$$

$$K_{\text{PrHNO}_2} = 4,6 \cdot 10^{-5}$$

Решение:

полн. самовытесн.

$$C_{\text{HNO}_2} = 0,1 \cdot \frac{10}{20} = 0,05 \text{ моль/л}$$

$$C_{\text{PrHNO}_2} = 0,1 \cdot \frac{10}{20} = 0,05 \text{ моль/л}$$

Для диссоциирующей кислоты:

$$C_{\text{H}^+} = C_{\text{HNO}_2} = 0,05 \text{ моль/л}$$

$$C_{\text{PrH}^+} = C_{\text{PrHNO}_2} = 0,05 \text{ моль/л}$$

$$[C_{\text{H}^+}][C_{\text{PrH}^+}] = 4,6 \cdot 10^{-5} \cdot 10^{-2} = 4,6 \cdot 10^{-7}$$

$$pH = -\lg [H^+] = -\lg 1,72 \cdot 10^{-12} = 12 - \lg 1,72 = 11,76$$

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microprocessor shifts the control from the main program to process the incoming request after the request is completed the

control goes back to the main program interrupt signals present in 8085 are intr