A. Balance each of the following chemical equations.

1. Fe + O, 
$$\rightarrow$$
 Fe,O,

$$2. \quad C_{1}H_{4} + C_{2}O_{2} \rightarrow C_{2}O_{2} + H_{2}O_{3}$$

4. 
$$NH_3 + O_2 \rightarrow NO + H_2O$$

5. 
$$\underline{\text{HClO}}_{\downarrow} + \underline{\text{P}}_{\downarrow}\text{O}_{\downarrow} \rightarrow \underline{\text{H}}_{\downarrow}\text{PO}_{\downarrow} + \underline{\text{H}}_{\downarrow}$$

$$\mathbf{6.} \quad \mathbf{S_s} \quad + \quad \mathbf{O_2} \quad \rightarrow \quad \mathbf{SO_3}$$

7. 
$$\underline{\text{HNO}}_{\text{NO}}$$
 +  $\underline{\text{NaHCO}}_{\text{NO}}$   $\rightarrow$   $\underline{\text{NaNO}}_{\text{NO}}$  +  $\underline{\text{NaNO}}_{\text{NO}}$ 

8. 
$$CO$$
, +  $H$ , O  $\rightarrow$   $C_6H_0O_6$  +  $O$ ,

9. 
$$Z_n$$
 +  $A_gNO_s$   $\rightarrow$   $Z_n(NO_s)_s$  +  $Z_n$ 

- B. Solve the following problems and write your solutions and answers in the space provided for. Use the atomic weights in the periodic table of your procedure manual.
  - Iron (II) sulfate, FeSO<sub>4</sub>, is prescribed for the treatment of anemia.
    - a) How many moles of FeSO4 are present in a

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