

# ACIDS, BASES AND SALTS

|                                                                                                                                                |                                                                                                                                                              |
|------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Definition of an acid</b><br>An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions | <b>What causes the acidic properties of acids?</b><br>The hydrogen ions ( $H^+$ ) cause the acidic properties and these are formed in the presence of water. |
|------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------|

|                                                                                    |                                                                                                                                                                |
|------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Another term to refer to an acid</b><br>An acid is called a <u>proton donor</u> | <b>Why an acid is also called a proton donor?</b><br>It's because an acid provides protons or hydrogen ions ( $H^+$ ) to other substances during the reaction. |
|------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------|

|                                                            |                                                 |                                                                                                       |                                                                                                                 |
|------------------------------------------------------------|-------------------------------------------------|-------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------|
| <b>Substance to which an acid provides protons</b><br>BASE | <b>A base is then called</b><br>PROTON ACCEPTOR | <b>Why the base is called a proton acceptor?</b><br>Because the base accepts hydrogen ions from acids | <b>Equation for the reaction between acid and base</b><br>$H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$ |
|------------------------------------------------------------|-------------------------------------------------|-------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------|

| Common laboratory acids |           | These three common laboratory acids are also called | Why these acids are also called mineral acids?                                                              |
|-------------------------|-----------|-----------------------------------------------------|-------------------------------------------------------------------------------------------------------------|
| Hydrochloric acid       | HCl       | MINERAL ACIDS                                       | They are derived from mineral salts i.e chlorides for HCl, sulphates for $H_2SO_4$ and nitrates for $HNO_3$ |
| Sulphuric acid          | $H_2SO_4$ |                                                     |                                                                                                             |
| Nitric acid             | $HNO_3$   |                                                     |                                                                                                             |

| Other mineral acids known |           | Mineral salts from which the acid is derived | Organic acids known          | Naturally occurring acids               |
|---------------------------|-----------|----------------------------------------------|------------------------------|-----------------------------------------|
| Sulphurous acid           | $H_2SO_3$ | Derived from sulphites                       | Ethanoic acid ( $CH_3COOH$ ) | CITRIC ACID from lemons                 |
| Carbonic acid             | $H_2CO_3$ | Derived from carbonates                      |                              | TARTARIC ACID from grapes               |
| Phosphoric acid           | $H_3PO_4$ | Derived from phosphates                      | Methanoic acid ( $HCOOH$ )   | ACETIC ACID from vinegar                |
| Nitrous acid              | $HNO_2$   | Derived from nitrites                        |                              | LACTIC ACID from sour milk              |
|                           |           |                                              |                              | Hydrochloric acid from digestive juices |

|                                                                             |                                                                                                             |                                                                                                                                                                               |
|-----------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Whenever an acid is dissolved in water, it produces</b><br>HYDROGEN IONS | <b>Term given to the number of hydrogen ions produced by one molecule of an acid</b><br>BASICITY OF AN ACID | <b>Definition of basicity of an acid</b><br>BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> . |
|-----------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|

# Acids And Bases Study Guide

**David R. Klein, Laurie S. Starkey**



## **Acids And Bases Study Guide:**

## **Acids And Bases Study Guide** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the energy of words has be much more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Acids And Bases Study Guide**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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