

Name : _____ Date : _____

Acids and Bases Problems

- Without doing any calculations, identify each of the following solutions as acidic or basic:
 - $[\text{H}_3\text{O}^+] = 3.4 \times 10^{-8} \text{ M}$ _____
 - $[\text{OH}^-] = 1.8 \times 10^{-5} \text{ M}$ _____
 - $[\text{OH}^-] = 6.5 \times 10^{-12} \text{ M}$ _____
 - $[\text{H}_3\text{O}^+] = 2.6 \times 10^{-4} \text{ M}$ _____
- A solution of HBr has $[\text{H}_3\text{O}^+] = 4.5 \times 10^{-3} \text{ M}$. Calculate the pH of this solution.

- Complete the following table.

pH	$[\text{H}_3\text{O}^+]$	$[\text{OH}^-]$	Acidic, basic, or neutral?
	5.4×10^{-4}		
	1.3×10^{-5}		
10.75	1.8×10^{-11}	5.6×10^{-4}	

- If the pH of a solution is 10.3, what is the $[\text{H}^+]$ concentration?
- If the $[\text{H}^+]$ is $2.1 \times 10^{-12} \text{ M}$ HClO_4 , what is the pH? Is the solution ACIDIC, BASIC, or NEUTRAL?
- Calculate the pH if the $[\text{OH}^-]$ concentration is $5.9 \times 10^{-1} \text{ M}$? Is the solution ACIDIC, BASIC, or NEUTRAL?

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