

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

Crystalline Solids

- Have a regular, repeating arrangement of particles (atoms, ions, or molecules).
- Have sharp melting points.
- Can be cleaved along flat planes.

Amorphous Solids

- Do not have a regular, repeating arrangement of particles.
- Do not have sharp melting points.
- Do not cleave along flat planes.

Intermolecular Forces

- London Dispersion Forces:** Present in all molecules, but only significant in nonpolar molecules.
- Dipole-Dipole Forces:** Present in polar molecules.
- Hydrogen Bonding:** A special type of dipole-dipole force between molecules containing N-H, O-H, or F-H bonds.

Periodic Table & Ionization Energies

Ionization Energy Trends:

- Increases from bottom-left to top-right.
- Decreases from top-left to bottom-right.

Properties of the Periodic Table

- Metals:** Located on the left side of the table.
- Nonmetals:** Located on the right side of the table.
- Metalloids:** Located between metals and nonmetals.
- Groups:** Vertical columns of elements.
- Periods:** Horizontal rows of elements.

Electronegativity Table

Electronegativity Trends:

- Increases from bottom-left to top-right.
- Decreases from top-left to bottom-right.

Types of Chemical Calculations

Stoichiometry: Calculations involving the relative amounts of reactants and products in a chemical reaction.

Thermochemistry: Calculations involving heat and energy changes in chemical reactions.

Equilibrium: Calculations involving the concentrations of reactants and products at equilibrium.

Thermodynamics

First Law of Thermodynamics: Energy cannot be created or destroyed, only transferred or converted.

Second Law of Thermodynamics: The entropy of the universe always increases.

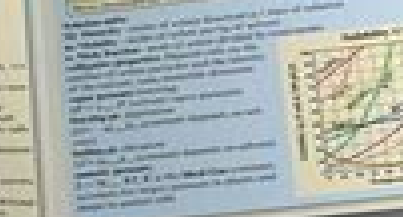
Third Law of Thermodynamics: The entropy of a perfect crystal is zero at absolute zero.

Equilibrium & Solubility

Equilibrium Constants: K_c (concentration), K_p (pressure), K_a (acid dissociation), K_b (base dissociation).

Solubility Rules:

- All nitrates, acetates, and most salts of the alkali metals are soluble.
- All chlorides, bromides, and iodides are soluble, except for those of Ag, Hg, and Pb.
- All sulfates are soluble, except for those of Ba, Sr, and Pb.



Acids & Bases

Arrhenius Definition: Acids produce H^+ ions in water; bases produce OH^- ions in water.

Brønsted-Lowry Definition: Acids are proton donors; bases are proton acceptors.

Conjugate Acid-Base Pairs: Differ by one proton.

Redox Reactions

Oxidation: Loss of electrons.

Reduction: Gain of electrons.

Oxidation States: Assigned to atoms in a molecule to track electron transfer.

Half-Reactions: Separate oxidation and reduction processes.

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Wenbin Ji



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