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Student Exploration: Bohr Model of Hydrogen

Vocabulary: absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

[Note to teachers and students: This Gizmo was designed as a follow-up to the Bohr Model: Introduction Gizmo. We recommend doing that activity before trying this one.]

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- What happens to an electron when it absorbs a tiny packet of light called a **photon**?
the electron absorbs the energy from it, causing a higher energy level
- What happens to an electron when it emits a photon?
the electrons lose energy, causing a lower energy level

Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the **absorption spectrum** of hydrogen. An **emission spectrum** is created when hydrogen gas emits light.

The **Bohr model** of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the *Bohr Model of Hydrogen* Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

- With the **Laser energy** set to 7.0 eV, click **Play** (). Observe the blue electron on the ORBITALS pane. Do 7.0-eV photons have any effect on the electron? No, it doesn't, the electron stays at lvl 1.
- Click **Pause** (). Set the **Laser energy** to 12.1 eV and click **Play**. What happens to the electron now? the electron shakes up and down and moves up by 2 orbital (level 3)
- Under **Go to energy level**, select **1**. Notice that a photon is emitted by the electron. What is the energy of the emitted photon? (12.1 eV)

Bohr Model Of Hydrogen Gizmo Answer Sheet

Ricardo Simeoni, Griffith University



Bohr Model Of Hydrogen Gizmo Answer Sheet:

Bohr's Model of Atomic Hydrogen Extended to Include Electron Rotational Kinetic Energy Ricardo Simeoni, Griffith University, 2003

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