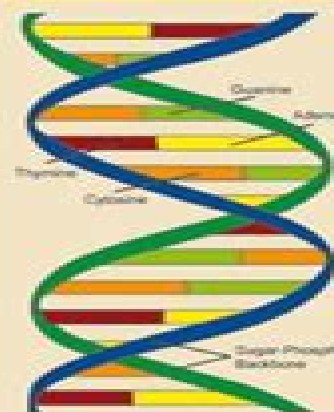


BIOCHEMISTRY

BIOCHEMICAL PERIODIC TABLE

Key Elements in the Body																																																																																					
1 H	3 Li	11 Na	12 Mg	19 K	20 Ca	21 Ti	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Al	32 Si	33 P	34 S	35 Cl	36 Ar	37 K	38 Ca	39 Sc	40 Ti	41 V	42 Cr	43 Mn	44 Fe	45 Co	46 Ni	47 Cu	48 Zn	49 Ga	50 Ge	51 As	52 Se	53 Br	54 Kr	55 Rb	56 Sr	57 Y	58 Zr	59 Nb	60 Mo	61 Tc	62 Ru	63 Rh	64 Pd	65 Ag	66 Cd	67 In	68 Sn	69 Sb	70 Te	71 I	72 Xe	73 La	74 Ce	75 Pr	76 Nd	77 Pm	78 Sm	79 Eu	80 Gd	81 Tb	82 Dy	83 Ho	84 Er	85 Tm	86 Yb	87 Fr	88 Ra	89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

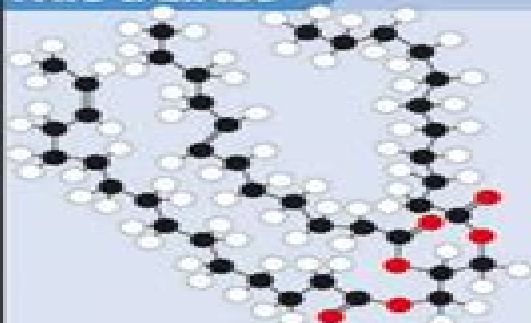
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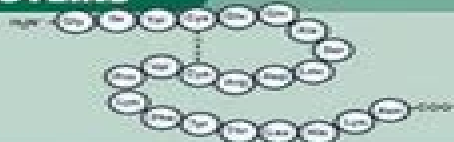
CARBOHYDRATES



FATS & LIPIDS



PROTEINS



BROADER CHEMICAL PRINCIPLES

A. Intermolecular Forces

1. **Electrostatic:** Strong interaction between ions; for charges q_1 and q_2 , separated by r_{12} , and solvent dielectric constant, ϵ ; water has large ϵ ; stabilizes transition.

2. **Posterior Distribution:** $\pi(\theta | y)$

3. **Dipole moment, μ :** Asymmetric electron distribution gives partial charge to

Abstract

3. **London forces (dispersion):** Attraction due to induced dipole moments; force increases with μ
5. **Dipole-dipole interactions:** The positive end of one dipole is attracted to the negative end of another dipole; strength increases with μ
6. **Hydrogen bonding:** Enhanced dipole interaction between bonded H and the lone-pair of neighboring O, N or S; gives "structure" to liquid water; solubilizes alcohols, fatty acids, amines, sugars, and amino acids
- 

B. Types of Chemical Groups

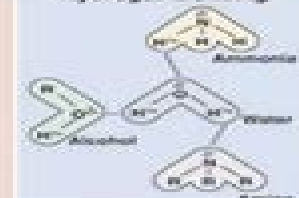
1. **Hydrophobic** = **Lipophilic**: Repelled by polar group; insoluble in water; affinity for nonpolar
EX: alkane, arene, silicone
2. **Hydrophilic** = **Lipophobic**: Affinity for polar group; soluble in water, repelled by nonpolar
EX: alcohol, amine, carboxylic acid

Energy = $\frac{1}{2} \frac{q_1 q_2}{r_{12}}$

Richard J. Stroh



Hydrogen Bonding



Osmotic Pressure
 $\Pi = \text{IMRT}$

3. **Amphipathic:** Polar and nonpolar functionality, common for most biochemical molecules: fatty acids, amino acids and nucleotides

C. Behavior of Solutions

1. **Miscible:** 2 or more substances form 1 phase; occurs for polar + polar or nonpolar + nonpolar
2. **Immiscible:** 2 liquids form aqueous and organic layers; compounds are partitioned between the layers based on chemical properties (acid/base, polar, nonpolar, ionic)

1. Physiological principles:

- b. **Biochemical example – Osmotic Pressure:** Water diffuses through a semi-permeable membrane from a hypotonic to a hypertonic region; the flow produces a force, the osmotic pressure, on the hypertonic side.

H: Osmotic pressure (in atm)
 i : Van't Hoff Factor = # of ions per solute molecule
 M: Solution molarity (moles/L)
 R: Gas constant = 0.082 L atm mol⁻¹ K⁻¹
 T: Absolute temperature (in Kelvin)

4. *Collective and self groups*

- a. Henry's Law: The amount of gas dissolved in a liquid is proportional to the partial pressure of the gas
- b. Carbon dioxide dissolves in water to form carbonic acid
- c. Oxygen is carried by hemoglobin in the blood
- d. Pollutants and toxins dissolve in bodily fluids; react with tissue and interfere with reactions

EX: Sulfur oxides and nitrogen oxides yield acids, ozone oxidizes lung tissue, hydrogen cyanide disables the oxidation of glucose

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Denise Guinn



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