

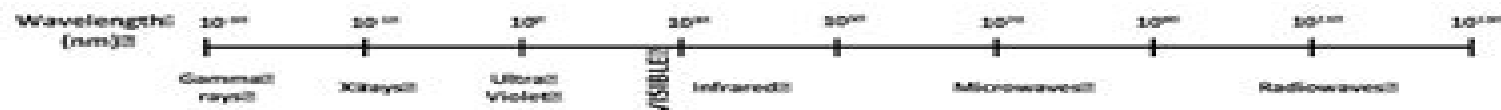
General Chemistry 1: Exam Appendix

Constants	Equations		
$R = 8.314 \text{ J/mol}\cdot\text{K}$ $R = 0.082 \text{ L}\cdot\text{atm/mol}\cdot\text{K}$ $h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$ $c = 3.00 \times 10^8 \text{ m/s}$ $N_A = 6.022 \times 10^{23} \text{ particles/mole}$	$E = (-2.18 \times 10^{-18} \text{ J}) Z^2/n^2$ $KE = (1/2)mv^2$ $PV = nRT$	$E = hc/\lambda$ $KE = (3/2)nRT$ $T(\text{K}) = T(^{\circ}\text{C}) + 273$	$E = mc^2$

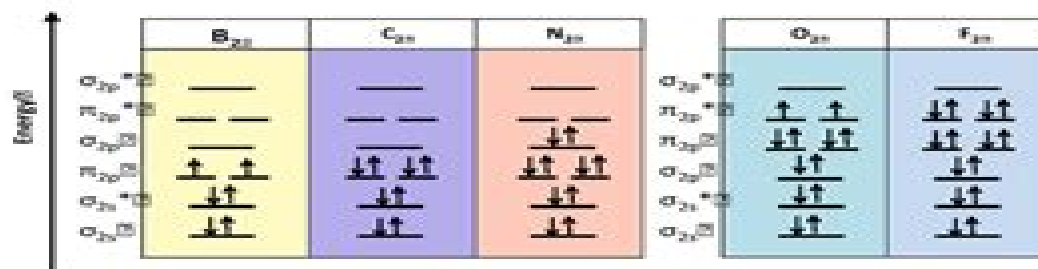
Solubility Rules

- All ammonium and group 1A salts are soluble
- All nitrates are soluble
- All acetates, chlorates, and perchlorates are soluble
- Most chlorides, bromides, and iodides are soluble - except silver, mercury, and lead
- Most sulfates are soluble - except calcium, strontium, barium, and lead
- Phosphates are insoluble (except ammonium and Group 1A)
- Carbonates are insoluble (except ammonium and Group 1A)
- Chromates (CrO_4^{2-}) are insoluble (except ammonium, Group 1A, and magnesium)
- Hydroxides are insoluble (except ammonium, Group 1A, barium, strontium, calcium)
- Sulfides are insoluble (except ammonium, Group 1A, magnesium, calcium, strontium, barium)

Electromagnetic Spectrum



Molecular Orbital Arrangements



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CL Gary



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