

Biology 160 Exam 1 Study Guide

L1: Organization & Function of the Cell

5 Big Ideas:

1. **Evolution** – populations of organisms and cells, including their molecular components, that have changed over time through both selective and non-selective evolutionary processes.
2. **Structure & Function** – all living systems (cells, organisms, ecosystems, etc) are made of structural components whose arrangement determines the function of the systems.
3. **Information Flow & Storage** – information and signals are used and exchanged within and among cells and organisms to direct their functioning.
4. **Transformations of Energy and Matter** – all living things acquire, use, and release and cycle matter and energy for cellular / organismal functioning.
 - a. Life obeys the laws of physics and chemistry
 - b. All living things require input of energy and matter
 - c. Energy can be neither created nor destroyed, but can change form
5. **Systems** – living systems are interconnected, and they interact and influence each other on multiple levels.

L2: Chemical Elements & Chemical Bonds

- **Proton** – positively charged atomic particle located within the nucleus of an atom with a charge of 1+ and a mass of approximately 1 Da.
- **Electron** – negatively charged atomic particle located outside of the nucleus in the orbitals of the atom with a charge of 1- and a mass of approximately 0 Da.
- **Neutron** – a neutrally charged atomic particle located within the nucleus of the atom with no relative charge and a mass of approximately 1 Da.
- **Atomic Number** – the total # of protons in the nucleus of an atom.
- **Mass Number** – the total # of protons + neutrons in the nucleus of an atom.
- **Isotopes** – two atoms with the same # of protons but with DIFFERENT # of neutrons are considered to be isotopes of the same element.
 - o Electrons move around the nuclei in specific regions called orbitals & each orbital can hold up to 2 electrons.
- **Valence shell** – the outermost shell of electrons.
- **Valence** – the # of unpaired electrons in the valence shell of an atom as well as the number of bonds that atom is likely to form. (carbon = 4, nitrogen = 3)
 - o Unpaired electrons are relatively uncommon in nature & atoms that carry unpaired electrons are usually rather reactive.
 - o Unpaired electrons in the valence shell can form covalent bonds.
 - Ex – 2 hydrogen atoms (each with one unpaired electron) can form a covalent bond with 2 shared electrons.

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Mcdougal Littell



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