



Asus Me172v Manual

Y Pai



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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C_3H_8 what is the percent yield $3CO + 7H_2 \rightarrow C_3H_8 + 3H_2O$

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web a write a balanced equation for the reaction $Al + O_2 \rightarrow Al_2O_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g } Al \times \frac{1 \text{ mol } Al}{27.0 \text{ g } Al} \times \frac{2 \text{ mol } Al_2O_3}{4 \text{ mol } Al} \times 101.96 \text{ g } Al_2O_3 = 120.1 \text{ g } Al_2O_3$ c determine the percent yield $\frac{101.96 \text{ g } Al_2O_3}{120.1 \text{ g } Al_2O_3} \times 100 = 84.9\%$

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