

Chemical Equilibrium Practice Problems

1. Oxides of nitrogen form from air heated to high temperatures inside auto engines. What is the equilibrium concentration of nitrogen monoxide if the equilibrium concentration of nitrogen is 0.36 M and that of oxygen is 0.010 M?

$$K = 0.0025$$



2. A flask is initially 0.050 M in both NO_2 and N_2O_4 . Which will be present in greater amount at equilibrium?

$$K = 0.50$$



3. Carbon dioxide forms from the reaction of calcium carbonate with an acid.



Analyze the changes below for the effectiveness of increasing the amount of carbon dioxide at equilibrium

a. add Salt, NaCl

b. add more acid

c. add a base

4. Chlorine for pool disinfection is often formed by mixing solid bleach and hydrochloric acid:



Analyze the changes below for effectiveness in increasing the amount of Cl_2 concentration at equilibrium

a. add salt, NaCl

b. add a base

c. add more solid bleach.

5. Catalytic converters change harmful carbon monoxide into safer carbon dioxide:



If this reaction was done in a smaller reaction chamber, would the equilibrium shift right or left?

6. Ammonia gas is formed through the Haber Process:



If this reaction was done in a larger reaction chamber, would the equilibrium shift right or left?

7. Chloroform, CHCl_3 , can be converted to carbon tetrachloride, CCl_4 , by the reaction below:



If you begin with 0.10 M of all substances, will the reaction form more products or more reactants?

8. In a 10.0 L vessel, 0.500 mol H_2 and 0.500 mol I_2 react at 448°C according to the equation



At this temperature $K = 50$ for this reaction. Find the equilibrium concentration for all three species.

Aqueous Chemical Equilibrium Practice Problems

R Pring



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