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Homer L. Patterson



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 $3\text{Ca(OH)}_2 + 1\text{Al}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{Al(OH)}_3$ type of reaction double displacement
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 $2\text{Fe} + 3\text{MgO} \rightarrow 2\text{FeO} + 3\text{Mg}$ type of reaction double displacement
 $4\text{HCl} + 1\text{Ca(OH)}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{CaCl}_2$ type of reaction double displacement

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 $\text{NaOH} + \text{KNO}_3 \rightarrow \text{NaNO}_3 + \text{KOH}$ type of reaction double displacement
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ type of reaction synthesis
 $2\text{Fe} + 6\text{NaBr} \rightarrow 2\text{FeBr}_3 + 6\text{Na}$ type of reaction single displacement
 $\text{MgSO}_4 + 2\text{NH}_4\text{OH} \rightarrow \text{Mg(OH)}_2 + (\text{NH}_4)_2\text{SO}_4$ type of reaction double displacement
 $2\text{HCl} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCl}$ type of reaction double displacement

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 $2\text{Fe} + 3\text{MgO} \rightarrow 2\text{FeO} + 3\text{Mg}$ type of reaction double displacement
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 $\text{PbSO}_4 + \text{PbSO}_3 + \text{O}_2 \rightarrow$ type of reaction 6 $\text{NH}_3 + \text{I}_2 \rightarrow$