



Atoms and molecules

an atom and its structure

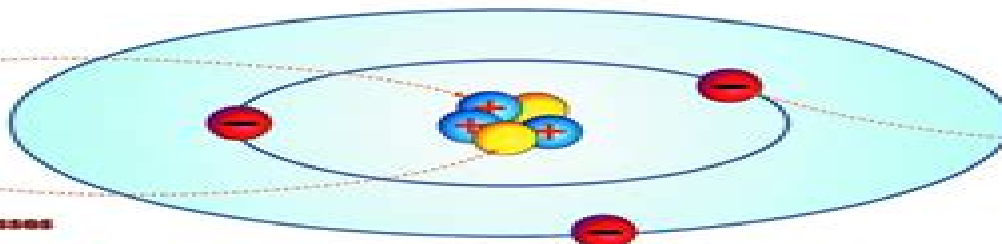
Atoms are made up of particles of matter known as **subatomic** or **fundamental particles**. The three main subatomic particles are called **protons**, **neutrons** and **electrons**. An atom consists mostly of empty space. Almost all of the atom's mass is concentrated in a structure at its centre called the **nucleus**. The nucleus consists of protons and neutrons grouped together. Electrons orbit the nucleus.

A **proton** is a positively charged particle

A **neutron** has no electric charge

Protons and neutrons have almost the same masses

An atom usually contains an equal number of positively charged protons and negatively charged electrons. This makes it electrically neutral.



An **electron** is a negatively charged particle. It is kept in orbit around the nucleus because it is attracted to the positively charged protons.

This attraction holds the atom together. An electron's mass is approximately 1/1840 times smaller than the mass of a proton or neutron.

electrons and their energy levels

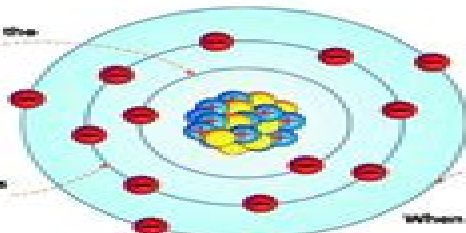
The way in which an atom reacts with other atoms (its **chemical properties**) is determined by the configuration of its electrons. Atoms can gain, lose or share electrons when they combine to form molecules. The electrons in an atom exist in layers called **energy levels** or **shells**. Electrons are arranged in the energy levels according to certain rules:

The first shell outside the nucleus holds up to two electrons

The second shell holds up to eight electrons

The third shell holds up to eight electrons.

When a shell is full, a new one is started.



Atomic Number	13
Atomic Symbol	Al
Atomic Mass	26.981 538

Number of electrons per shell (electron configuration)

The **atomic number** of an element tells you how many electrons that element has. An atom has as many energy levels as it needs to arrange its electrons. The number of electrons in the outer energy level determines how **reactive** an atom is. **Atoms like their energy levels to be full.** The most stable atoms are those with full outer energy levels, like helium. The most reactive atoms, like sodium, are those that only need to lose or gain one electron to obtain a full outer shell.

differences between mixtures & compounds

MIXTURES (e.g. iron and sulphur)

- "pure substances"
- "the proportions of the parts can vary (changed)"
- "the components can be elements or compounds"
- "the components are loosely put together"
- "the properties of each part do not change"
- "the properties of the mixture are the same as the properties of all its parts."

COMPOUNDS (e.g. iron sulphide)

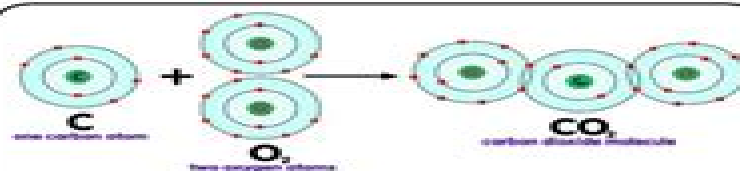
- "pure substances"
- "the proportions of the parts are constant"
- "the components are elements"
- "the components are held together by chemical bonds."
- "the properties of the compound are different from the properties of the elements forming it, and from a mixture of these elements."

The mixture formed when one substance **dissolves** into another is called a **solution**. Sugar that is dissolved in water is a solution. There are always two parts of a solution: the **solute** and the **solvent**. The component which forms the greatest part of the solution is the solvent. **Water** is the **solvent** of the sugar solution. The **sugar** is the **solute**.

molecules

While an atom is the smallest particle of an element, **molecules** are the smallest particles of a compound. A molecule is formed when at least two atoms combine. These may be of the same element, or of different elements. If the atoms in a molecule are of different elements, it is called a compound.

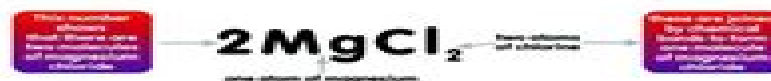
An example of such a compound is the carbon dioxide molecule:



A carbon atom has 4 electrons on its outer shell, while an oxygen atom has 6 electrons on its outer shell. It is thus necessary for 2 oxygen atoms to bond with 1 carbon atom so that the 2 oxygen atoms can fill their outer shells with 2 electrons, respectively, from the carbon atom ensuring that both oxygen atoms have full outer shells (8 required for 2nd shell) as does the carbon atom (2 required for 1st shell).

compounds & formulae

Compounds are formed when different elements combine. Like elements, each compound has a name and a symbol. The symbol of a compound is made of the symbols of the elements which form it, and is called a **formula**. A formula is used to show the atoms in a molecule. The numbers in a formula give the number of atoms of each kind of element in the molecule. The number of atoms of each element in that molecule is written after the symbol of the element, in small letters below the line e.g. O^2 which is a molecule of oxygen. A number written immediately before the formula of a molecule tells you how many molecules there are.



Atoms And Molecules A Childs Guide To Chemistry

Scott C. Dulebohn



Atoms And Molecules A Childs Guide To Chemistry:

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three principles that all aspects of chemical education should be associated with research that the development of opportunities for chemical education should be both a continuous process and be linked to research and that the professional development of all those associated with chemical education should make extensive and diverse use of that research It is intended for pre service and practising chemistry teachers and lecturers chemistry teacher educators chemical education researchers the designers and managers of formal chemical curricula informal chemical educators authors of textbooks and curriculum support materials practising chemists and chemical technologists It addresses the relation between chemistry and chemical education curricula for chemical education teaching and learning about chemical compounds and chemical change the development of teachers the development of chemical education as a field of enquiry This is mainly done in respect of the full range of formal education contexts schools universities vocational colleges but also in respect of informal education contexts books science centres and museums

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Table of Contents Atoms And Molecules A Childs Guide To Chemistry

1. Understanding the eBook Atoms And Molecules A Childs Guide To Chemistry
 - The Rise of Digital Reading Atoms And Molecules A Childs Guide To Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Atoms And Molecules A Childs Guide To Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atoms And Molecules A Childs Guide To Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atoms And Molecules A Childs Guide To Chemistry
 - Personalized Recommendations
 - Atoms And Molecules A Childs Guide To Chemistry User Reviews and Ratings
 - Atoms And Molecules A Childs Guide To Chemistry and Bestseller Lists

5. Accessing Atoms And Molecules A Childs Guide To Chemistry Free and Paid eBooks
 - Atoms And Molecules A Childs Guide To Chemistry Public Domain eBooks
 - Atoms And Molecules A Childs Guide To Chemistry eBook Subscription Services
 - Atoms And Molecules A Childs Guide To Chemistry Budget-Friendly Options
6. Navigating Atoms And Molecules A Childs Guide To Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Atoms And Molecules A Childs Guide To Chemistry Compatibility with Devices
 - Atoms And Molecules A Childs Guide To Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Atoms And Molecules A Childs Guide To Chemistry
 - Highlighting and Note-Taking Atoms And Molecules A Childs Guide To Chemistry
 - Interactive Elements Atoms And Molecules A Childs Guide To Chemistry
8. Staying Engaged with Atoms And Molecules A Childs Guide To Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atoms And Molecules A Childs Guide To Chemistry
9. Balancing eBooks and Physical Books Atoms And Molecules A Childs Guide To Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atoms And Molecules A Childs Guide To Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Atoms And Molecules A Childs Guide To Chemistry
 - Setting Reading Goals Atoms And Molecules A Childs Guide To Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Atoms And Molecules A Childs Guide To Chemistry
 - Fact-Checking eBook Content of Atoms And Molecules A Childs Guide To Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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