

4

Atomic Structure
Reviewsheet

A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are $^{32}_{16}\text{S}$ and $^{34}_{16}\text{S}$. An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 4.1
2. sub atomic 4.2
3. negative 4.3
4. protons 4.2
5. neutrons 4.2
6. nucleus 4.3
7. Rutherford 4.3
8. positive 4.3
9. electrons 4.3
10. protons (elec) 4.4
11. number 4.4
12. electrons 4.4
13. protons (elec) 4.5
14. neutrons 4.6
15. isotopes 4.6
16. atomic mass 4.7
17. 16 4.5
18. 16 4.6
19. 16 4.6
20. 18 4.6

Atomic Structure Answer Key Study Guide

Jiyuan Zhang



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Table of Contents Atomic Structure Answer Key Study Guide

1. Understanding the eBook Atomic Structure Answer Key Study Guide
 - The Rise of Digital Reading Atomic Structure Answer Key Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Answer Key Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atomic Structure Answer Key Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Answer Key Study Guide

- Personalized Recommendations
- Atomic Structure Answer Key Study Guide User Reviews and Ratings
- Atomic Structure Answer Key Study Guide and Bestseller Lists
- 5. Accessing Atomic Structure Answer Key Study Guide Free and Paid eBooks
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- 6. Navigating Atomic Structure Answer Key Study Guide eBook Formats
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 - Highlighting and Note-Taking Atomic Structure Answer Key Study Guide
 - Interactive Elements Atomic Structure Answer Key Study Guide
- 8. Staying Engaged with Atomic Structure Answer Key Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atomic Structure Answer Key Study Guide
- 9. Balancing eBooks and Physical Books Atomic Structure Answer Key Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atomic Structure Answer Key Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Atomic Structure Answer Key Study Guide
 - Setting Reading Goals Atomic Structure Answer Key Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Atomic Structure Answer Key Study Guide

- Fact-Checking eBook Content of Atomic Structure Answer Key Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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