

Flame Test Lab

Name Lani Simmons Period 8th Date 10/27

Objective: To recognize that the color of light emitted by a heated element is unique and can be translated into energy that is coming from the movement of electrons within the atoms.

Background: The natural position of an electron in an energy level is called its *ground state*. When an atom absorbs energy, it is in its *excited state*. This happens when electrons are heated in a flame or electrocuted. These excited electrons jump to higher energy levels. Eventually, the excited electrons fall back down to their ground state, releasing their extra energy as a packet of light called a *photon*.

The color of light corresponds to a specific amount of energy. This color varies from one substance to another, because the location of electrons within each substance is different. Just as we have unique fingerprints, each element has a unique color of light (amount of energy) that it releases. As a result, the colors we observe when a substance is heated may be used to identify the element.

Pre-lab Questions:

1. Fill in the blanks with the terms "*excited*, *ground*, or *light*."

When an atom *absorbs* energy, the electrons move from their ground state to their excited state. Then, the electrons *release* energy, as they fall back down from their excited state to their ground state. This release of energy is in the form of light.

2. What is a photon? A photon is a tiny particle that comprises waves of electromagnetic radiation. As shown by Maxwell, photons are just electric fields traveling through space. Photons have no charge, no resting mass, and travel at the speed of light.

Data: Hint: Each color will only be used once in the data table below.

Metal Salt	Circle the Color of the Flame that You Observed:					
Calcium Chloride	Yellow/Green	Yellow	Blue-green	Orange	Red	Crimson (Darker Red)
Copper Chloride	Yellow/Green	Yellow	Blue-green	Orange	Red	Crimson (Darker Red)
Lithium Chloride	Yellow/Green	Yellow	Blue-green	Orange	Red	Crimson (Darker Red)
Sodium Chloride	Yellow/Green	Yellow	Blue-green	Orange	Red	Crimson (Darker Red)
Barium Chloride	Yellow/Green	Yellow	Blue-green	Orange	Red	Crimson (Darker Red)
Strontium Chloride	Yellow/Green	Yellow	Blue-green	Orange	Red	Crimson (Darker Red)
Unknown	Yellow/Green	Yellow	Blue-green	Orange	Red	Crimson (Darker Red)

Unknown Number _____

Post-lab Questions:

1. In this lab, you tested six different metal salts. Each metal salt contained one metal atom and one chlorine atom. How do you know that the color observed for each salt came from exciting the *metal* atom and not the *chlorine* atom?

2. Identify your unknown *metal atom* in the space below. *CHLORIDE* should not be part of your answer!

Unknown Number	Identification

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