

Assigning Oxidation Numbers

Rule	Example
1. The oxidation number of any uncombined element is 0.	The oxidation number of $\text{Na}(s)$ is 0.
2. The oxidation number of a monatomic ion equals the charge on the ion.	The oxidation number of Cl^- is -1 .
3. The more-electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion.	The oxidation number of O in NO is -2 .
4. The oxidation number of fluorine in a compound is always -1 .	The oxidation number of F in LiF is -1 .
5. Oxygen has an oxidation number of -2 unless it is combined with F, in which it is $+1$ or $+2$, or it is in a peroxide, in which it is -1 .	The oxidation number of O in NO_2 is -2 .
6. Hydrogen's oxidation state in most of its compounds is $+1$ unless it is combined with a metal, in which case it is -1 .	The oxidation number of H in LiH is -1 .
7. In compounds, the elements of Groups 1 and 2 as well as aluminum have oxidation numbers of $+1$, $+2$, and $+3$, respectively.	The oxidation number of Ca in CaCO_3 is $+2$.
8. The sum of the oxidation numbers of all atoms in a neutral compound is 0.	The oxidation number of C in CaCO_3 is $+4$.
9. The sum of the oxidation numbers of all atoms in a polyatomic ion equals the charge of the ion.	The oxidation number of P in H_2PO_4^- is $+5$.

Assigning Oxidation Numbers Answers

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