

Additivity of Heats of Reaction: Hess's Law

In this experiment you will use a Styrofoam cup calorimeter to measure the heat released by three reactions/physical changes. One of them is the same as the sum of the other two. Therefore, according to Hess's Law of Constant Heat Summation (or just Hess's Law), the heat of reaction of the one reaction should be equal to the sum of the heats of reaction for the other two. The primary objective of this experiment is to confirm this law.

The reactions we will use in this experiment are:

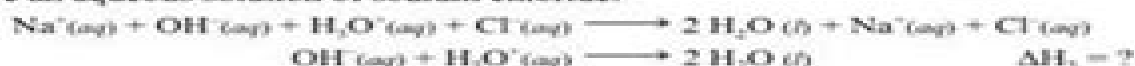
- (1) Solid sodium hydroxide dissolves in water to form an aqueous solution of ions.



- (2) Solid sodium hydroxide reacts with an aqueous solution of hydrochloric acid to form water and leave an aqueous solution of sodium chloride.



- (3) Aqueous solutions of sodium hydroxide and hydrochloric acid react to form water and leave an aqueous solution of sodium chloride.



You will use a Styrofoam cup (in a beaker) as a calorimeter, as shown in Figure 1. For purposes of this experiment, you may assume that the heat loss to the calorimeter and the surrounding air is negligible. Even if heat is lost to either of these it is a fairly constant factor in each part of the experiment and has little effect on the final results.



Figure 1

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