

EXPERIMENT 2

ELECTROCHEMISTRY: ELECTROCHEMICAL CELL AND THERMODYNAMIC FUNCTIONS

Objectives

1. To construct electrochemical cells
2. To determine the net cell potentials for three electrochemical cells
3. To calculate the net cell potential based on the potentials of the half-reactions that occur and to compare your experimental and calculated values.
4. To measure the Gibbs energy based on the cell potentials.

Introduction

The use of electrochemical cells to convert the Gibbs energy stored in the constituent half-reactions into electrical work is of enormous industrial as well as fundamental significance. We have all used batteries and these are simply galvanic cells that are similar to those you will be constructing in this experiment. In the laboratory, a typical electrochemical cell has the following general construction:

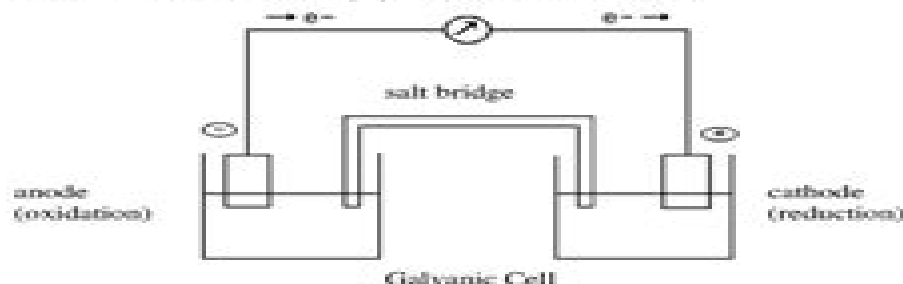


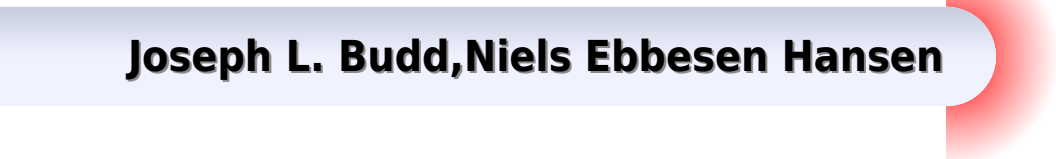
Figure 2.1: A schematic diagram of galvanic cell

The standard reduction potential is a quantitative measure of a substance's tendency to accept electrons under standard conditions (1 atm of gases and 1 M concentration) and is measured in volts. The more positive the reduction potential, the greater the tendency for reduction to occur. For the half-reaction that represents the oxidation of a metal, you can determine the standard oxidation potential (E°_{ox}) for the half-reaction. This value has the same numerical value as the standard reduction potential but the opposite sign. For example, aluminum ion has a standard reduction potential value of -1.65 V



Applied Thermodynamic Lab Student Manual

Joseph L. Budd, Niels Ebbesen Hansen



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