

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

Crystalline Solids

- Have a regular, repeating arrangement of particles (atoms, ions, or molecules).
- Have sharp melting points.
- Can be cleaved along flat planes.

Amorphous Solids

- Do not have a regular, repeating arrangement of particles.
- Do not have sharp melting points.
- Do not cleave along flat planes.

Intermolecular Forces

- London Dispersion Forces:** Present in all molecules, but only significant in nonpolar molecules.
- Dipole-Dipole Forces:** Present in polar molecules.
- Hydrogen Bonding:** Present in molecules containing N-H, O-H, or F-H bonds.

Periodic Table & Ionization Energies

Ionization Energy Trends:

- Increases from bottom-left to top-right.
- Decreases from top-left to bottom-right.

Properties of the Periodic Table

- Metals:** Located on the left side of the table.
- Nonmetals:** Located on the right side of the table.
- Metalloids:** Located between metals and nonmetals.
- Groups:** Vertical columns of elements.
- Periods:** Horizontal rows of elements.

Electronegativity Table

Electronegativity Trends:

- Increases from bottom-left to top-right.
- Decreases from top-left to bottom-right.

Types of Chemical Calculations

Stoichiometry

- Calculating the amount of product or reactant in a chemical reaction.

Thermochemistry

- Calculating the heat of reaction or the heat of formation.

Equilibrium

- Calculating the equilibrium constant or the concentration of a species at equilibrium.

Thermodynamics

First Law of Thermodynamics

$$\Delta E = q + w$$

Second Law of Thermodynamics

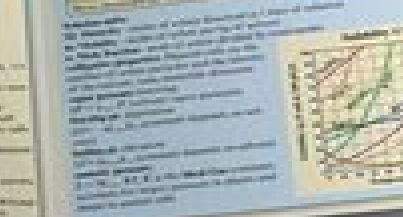
$$\Delta S_{univ} > 0$$

Third Law of Thermodynamics

$$\Delta S_{univ} = 0$$

Thermodynamic Data

Substance	ΔH_f° (kJ/mol)	ΔG_f° (kJ/mol)	S° (J/mol·K)
H ₂ (g)	0	0	130.7
O ₂ (g)	0	0	205.1
H ₂ O(l)	-285.8	-237.1	69.9
H ₂ O(g)	-241.8	-183.1	188.8



Equilibrium

Equilibrium Constant

$$K = \frac{[products]}{[reactants]}$$

Le Chatelier's Principle

If a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will shift in a direction that relieves the stress.

Acid-Base Equilibrium

Acid-Base Reactions

$$HA + B \rightleftharpoons A^- + BH^+$$

Acid-Base Equilibrium

$$K_a = \frac{[A^-][H^+]}{[HA]}$$

Acid-Base Titrations

Titration curves show the change in pH as an acid is titrated with a base. The curve starts at a low pH, rises sharply at the equivalence point, and then levels off at a high pH.

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