

# Acid-Base Titration

The diagram illustrates the four stages of an acid-base titration. A burette labeled 'Base of known concentration' is positioned above a flask containing 'Acid'. A pipette labeled 'Indicator' is used to add the indicator to the acid. Stage 1 shows the initial setup. Stage 2 shows the base being added drop by drop. Stage 3 shows the indicator changing color as the acid and base react. Stage 4 shows the solution turning a deep pink, indicating that the amount of base added has surpassed the amount of acid, which is a point to be avoided.

Base of known concentration

Indicator

Acid

1

2

3

4

A pH indicator is added to the flask containing an aqueous solution of the acid of unknown concentration.

The base is added slowly (drop by drop) to the acid and mixed repeatedly.

Equal amount of  $\text{OH}^-$  (base) and  $\text{H}^+$  (acid) causes indicator to change its color. At this stage we should stop titration.

If the amount of  $\text{OH}^-$  surpasses the amount of  $\text{H}^+$ , the solution changes color completely. This should be avoided.

# Ap Lab 14 Acid Base Titration

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