



Some Basic Concepts of Chemistry



$$\text{Velocity} = \frac{\text{Distance}}{\text{Time}}$$



$$\text{Acceleration} = \frac{\text{Velocity}}{\text{Time}}$$

$$1 \text{ amu} = 1.66056 \times 10^{-24} \text{ gram}$$

Atomic mass unit



$$\text{Density} = \frac{\text{Mass}}{\text{Volume}}$$



$$\text{Pressure} = \frac{\text{Force}}{\text{Area}}$$



$$^{\circ}\text{F} = \frac{9}{5} (^{\circ}\text{C}) + 32$$

$$\text{K} = ^{\circ}\text{C} + 273.15$$

$^{\circ}\text{F}$ - Fahrenheit
 $^{\circ}\text{C}$ - Celsius
 K - Kelvin



$$\text{Number of grams of atoms} = \frac{\text{Weight (in grams)}}{\text{gram atomic mass}}$$



$$\text{Molecular Mass } (A_x B C_y) = x \times [\text{Atomic mass of A}] + [\text{Atomic mass of B}] + y \times [\text{Atomic mass of C}]$$

X = Multiply

OR



$$\text{Molecular Mass} = \frac{\text{Actual mass of a molecule of matter}}{\frac{1}{12} \times \text{mass of one } ^{12}\text{C atom}}$$



$$\text{Gram Molecular No.} = \frac{\text{Mass (in grams)}}{\text{Molecular weight}}$$



$$\text{Equivalent weight of the element} = \frac{\text{Atomic weight of the element}}{\text{valency}}$$



$$\text{Equivalent weight of an element} = \frac{\text{formula weight of salt}}{\text{valency of the element in the salt}}$$



$$\text{Equivalent weight of an acid} = \frac{\text{Molecular weight of acid}}{\text{number of replaceable H}^+ \text{ ions}}$$



$$\text{Equivalent weight of a base} = \frac{\text{Molecular weight of base}}{\text{number of replaceable OH}^- \text{ ions}}$$



$$\text{Equivalent weight of a salt} = \frac{\text{Molecular weight of salt}}{\text{total positive or negative charge}}$$



$$\text{Equivalent weight of an oxidizing agent} = \frac{\text{Molecular weight of oxidizing agent}}{\text{number of electrons gained}}$$

$$\text{Equivalent weight of a reducing agent} = \frac{\text{Molecular weight of reducing agent}}{\text{number of electrons lost}}$$



$$\text{Equivalent weight of } \text{H}_2\text{O}_2$$

$\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
 Oxidation state of O in H_2O_2 is -1
 Oxidation state of O in H_2O is -2
 Change in oxidation state = 1



$$\text{Equivalent weight of } \text{H}_2\text{O}_2 = \frac{\text{Molecular weight of } \text{H}_2\text{O}_2}{2}$$

Equivalent weight of H_2O_2 is 17

Assessment Of Basic Chemistry Concepts Answer Sheet

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Assessment-Centered Teaching Kathryn DiRanna, 2008-03-07 Valuable for practitioners who wish to improve their teaching and their students learning and for researchers concerned with putting ideas of formative assessment into teaching practice Richard J Shavelson Margaret Jack Professor of Education Stanford University Drawing from conceptual principles and empirical findings that establish the crucial role of ongoing formative assessment the authors describe a professional development framework and program that prepares teachers to adjust their teaching to student thinking in the moment and to refine assessments to better reveal students understandings throughout instruction Joan I Heller Director Heller Research Associates A reflective process for integrating assessment and instruction Because assessment and instruction are two sides of the same coin it is critical for teachers to not only assess what students understand but also use that information to adjust their teaching Assessment Centered Teaching ACT is a unique practice that allows teachers to gather information during instruction to uncover learning gaps and guide students toward deeper understandings of complex ideas Suitable for all grade levels this resource describes how reflective practitioners can use the ACT portfolio to reflect on modify and improve their curriculum and instruction The forms included on the CD ROM guide teachers through the process This book encourages educators to Reflect on their own practice Systematically examine student progress toward identified learning goals over time Monitor instruction and assessment for continuous improvement Assessment Centered Teaching helps teachers professional developers and school teams establish quality instructional goals and implement ongoing formative assessment to promote student success

General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

Science Educator's Guide to Laboratory Assessment Rodney L. Doran, 2002 The book opens with an up to date discussion of assessment theory research and uses Then comes a wealth of sample assessment activities in biology chemistry physics and Earth science Keyed to the National Science Education Standards the activities include reproducible task sheets and scoring rubrics All are ideal for helping students reflect on their own learning during science lab

Resources in Education , 2000-04 **Resources for Teaching Elementary School Science** National Science Resources Center of the National Academy of Sciences and the Smithsonian Institution, 1996-03-28 What activities might a teacher use to help children explore the life cycle of butterflies What does a science teacher need to conduct a leaf safari for students Where can children safely enjoy hands on experience with life in an estuary Selecting resources to teach elementary school science can be confusing and difficult but few decisions have greater

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Focus on First Year Success Brenda Leibowitz, Antoinette Van der Merwe, Susan Van Schalkwyk, 2009-11-01 The importance of the first year experience is now well recognised This collection of papers makes a fascinating and important contribution to our understanding of students transition to higher education This is a scholarly engaging and illuminating text that is relevant not only in the context of South Africa but for anyone interested in student learning in the first year of university education David Gosling Plymouth University

Cehmistry Textbook for College and University USA Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA

Annual Catalogue United States Air Force Academy, 1983

United States Air Force Academy United States Air Force Academy,

Introduciton to educational measurement Victor Herbert Noll, 1957

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