

Aims

You have learned how to construct Born-Haber cycles and calculate lattice enthalpies from experimental data. In this activity you have the opportunity to apply your knowledge to some more challenging examples. You will need to have studied Topics 12.1 and 12.2 before attempting these problems.

Questions

- 1 a) Write an equation for the formation of aluminium oxide.
- b) Using the enthalpy change values in the table below, construct a Born-Haber cycle and calculate the lattice enthalpy of aluminium oxide.

Change	Enthalpy change	Value / kJ mol^{-1}
$2\text{Al}(s) + \frac{3}{2}\text{O}_2(g) \rightarrow \text{Al}_2\text{O}_3(s)$	Heat of formation	-1675.7
$\text{Al}(s) \rightarrow \text{Al}(g)$	Heat of atomisation	+324.3
$\text{Al}(g) \rightarrow \text{Al}^+(g)$	First ionisation energy	+578
$\text{Al}^+(g) \rightarrow \text{Al}^{2+}(g)$	Second ionisation energy	+1817
$\text{Al}^{2+}(g) \rightarrow \text{Al}^{3+}(g)$	Third ionisation energy	+2745
$\frac{1}{2}\text{O}_2(g) \rightarrow \text{O}(g)$	Heat of atomisation	+249.2
$\text{O}(g) \rightarrow \text{O}^-(g)$	First electron affinity	-141.4
$\text{O}^-(g) \rightarrow \text{O}^{2-}(g)$	Second electron affinity	+790.8

- c) Comment on the value you have calculated.
- 2 The standard enthalpy change of the formation of $\text{MgO} = -602\text{kJ mol}^{-1}$.

Calculate the standard enthalpy change for the reaction



Aqa Chemistry A2 Stretch And Challenge Answers

Alyn G. McFarland, Nora Henry



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Table of Contents Aqa Chemistry A2 Stretch And Challenge Answers

1. Understanding the eBook Aqa Chemistry A2 Stretch And Challenge Answers
 - The Rise of Digital Reading Aqa Chemistry A2 Stretch And Challenge Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Chemistry A2 Stretch And Challenge Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aqa Chemistry A2 Stretch And Challenge Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Chemistry A2 Stretch And Challenge Answers
 - Personalized Recommendations

- Aqa Chemistry A2 Stretch And Challenge Answers User Reviews and Ratings
- Aqa Chemistry A2 Stretch And Challenge Answers and Bestseller Lists
- 5. Accessing Aqa Chemistry A2 Stretch And Challenge Answers Free and Paid eBooks
 - Aqa Chemistry A2 Stretch And Challenge Answers Public Domain eBooks
 - Aqa Chemistry A2 Stretch And Challenge Answers eBook Subscription Services
 - Aqa Chemistry A2 Stretch And Challenge Answers Budget-Friendly Options
- 6. Navigating Aqa Chemistry A2 Stretch And Challenge Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Aqa Chemistry A2 Stretch And Challenge Answers Compatibility with Devices
 - Aqa Chemistry A2 Stretch And Challenge Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aqa Chemistry A2 Stretch And Challenge Answers
 - Highlighting and Note-Taking Aqa Chemistry A2 Stretch And Challenge Answers
 - Interactive Elements Aqa Chemistry A2 Stretch And Challenge Answers
- 8. Staying Engaged with Aqa Chemistry A2 Stretch And Challenge Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aqa Chemistry A2 Stretch And Challenge Answers
- 9. Balancing eBooks and Physical Books Aqa Chemistry A2 Stretch And Challenge Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aqa Chemistry A2 Stretch And Challenge Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aqa Chemistry A2 Stretch And Challenge Answers
 - Setting Reading Goals Aqa Chemistry A2 Stretch And Challenge Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aqa Chemistry A2 Stretch And Challenge Answers
 - Fact-Checking eBook Content of Aqa Chemistry A2 Stretch And Challenge Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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