

PHYSICAL CHEMISTRY



BASICS

Adiabatic process	Process in which no heat is transferred into or out of the system
Boiling point	Lowest temperature at which liquid under fixed pressure changes state from liquid to vapor (vapor pressure = pressure on liquid)
Carnot cycle	System undergoing a cyclical process involving two reversible isothermal processes and two reversible adiabatic processes
Carnot engine	The most efficient heat engine which follows the Carnot cycle
Chemical potential (μ)	Measure of a chemical system's tendency to change (undergo a reaction, form a new phase)
Critical point	Temperature where liquid and vapor are identical
Cyclical process	A process which repeats a series of changes in the system
Enthalpy (H)	The quantity of heat in a substance per unit mass under constant pressure
Entropy (S)	A measure of degree of randomness of a system
Equilibrium state	State in which state's macroscopic properties (such as temperature) are well defined and do not change with time
Heat (q or Q)	A form of energy; an energy transfer due to temperature difference between system and surroundings
Heat capacity (C)	Total amount of heat needed to produce a one degree rise in temperature of a given substance
Heat engine	Device that converts heat into mechanical energy

ENERGY CONVERSION FACTORS

1 J	1 J = 1 kg m ² /s ² = 0.239 cal = 10 ⁷ erg = 2.78 × 10 ⁻⁷ kWh = 9.481 × 10 ¹³ eV
1 cal	1 cal = 10 ⁻³ kWh = 3.968 × 10 ⁷ erg = 4.186 × 10 ⁷ erg = 4.185 J
1 kWh	3.6 × 10 ⁶ J = 8.601 × 10 ⁵ cal = 3413 BTU = 3.6 × 10 ¹¹ erg
1 BTU	1055 J = 252 cal = 2.93 × 10 ⁴ kWh = 1.055 × 10 ¹⁰ erg

PRESSURE CONVERSION FACTORS

1 atm	1.01325 × 10 ⁵ Pa = 1.01 × 10 ⁵ N/m ² = 1.01325 bar = 760 Torr = 760 mm Hg = 29.9 in Hg
1 Pa	9.869 × 10 ⁻⁶ atm = 7.5 × 10 ⁻⁶ mm Hg = 1 N/m ²
1 millibar	10 ² Pa

EQUILIBRIUM CONSTANTS

- For reaction, $aA + bB \rightleftharpoons cC + dD$

GAS EQUILIBRIUM

- For gases, the equilibrium constant is expressed in terms of partial pressures

$$K_p = \frac{[P_C]^c [P_D]^d}{[P_A]^a [P_B]^b}$$

$$K_p = K_c (RT)^{\Delta n}$$

where $\Delta n = (c + d) - (a + b)$

CHEMICAL EQUILIBRIUM

- At equilibrium, concentrations are at a fixed ratio

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Internal energy (U or E)

Irreversible process

Isobaric process

Isochoric process

Isothermal process

Mass

Melting point

Pressure (p or P)

Quantity of heat

Reversible process

State function (state variables)

Temperature (T)

Triple point

Work (w or W)

The total kinetic and potential energy at the molecular level

Non-equilibrium process in which no small change in conditions could make the process go in reverse

Process in which pressure remains constant

Process in which volume remains constant

Process in which temperature remains constant

Amount of matter in an object

The lowest temperature at which a solid under a fixed pressure changes state from a solid to a liquid

The force acting on a unit area

Total kinetic energy of substance's molecules

Process whose direction may be reversed by small change in one factor; system is always near equilibrium

Property of systems which has a definite value for each state, and does not depend on how the state is reached (such as temperature, pressure, volume)

Average kinetic energy of molecules of a substance (Note: has to be combined with quantity of heat)

Temperature and pressure at which 3 phases of a substance (solid, liquid, vapor) co-exist in equilibrium

Form of energy transfer that may be represented as a force acting through a distance

TEMPERATURE CONVERSION FACTORS

$$T_F = 9/5 \times T_C + 32$$

$$T_C = 5/9 (T_F - 32)$$

$$T_F = T_C + 459.67$$

$$T_K = T_C + 273.15$$

F = Fahrenheit; C = Celsius; K = Kelvin; R = Rankine

CONSTANTS

Avogadro Constant (N_A)	6.0221367 × 10 ²³ molecules/mol
Boltzmann Constant (k)	1.380658 × 10 ⁻²³ J/K
Stefan-Boltzmann Constant (σ)	5.67051 × 10 ⁻⁸ W/m ² /K ⁴
Molar Gas Constant (R)	8.314510 J/mol K
Molar Volume of an Ideal Gas (V_m)	0.082058 L atm/mol K
	22.41496 L/mol (273.15 K, 1 atm)

$$Q_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

- At start of a reaction, concentrations are at a ratio
- If $Q_c < K_c$, then the reaction proceeds from left to right
- If $Q_c = K_c$, then the system is at equilibrium
- If $Q_c > K_c$, then the reaction proceeds from right to left



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