

* Show that, the entropy remains constant during a reversible cyclic change.

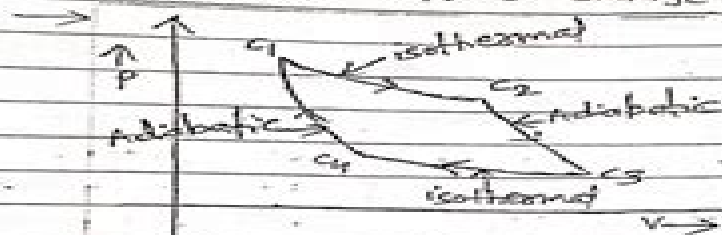


Fig: Carnot's cycle

Considered a reversible cycle C_1, C_2, C_3, C_4, C_1 as shown in fig. It consists of 4 operations
 1) From C_1 to C_2 heat Q_1 is absorbed by the system at temp. T_1 . The change in entropy of the system is S_1 . The change is positive.

2) From C_2 to C_3 , no heat is absorbed or rejected by the system, hence no change in entropy

3) From C_3 to C_4 , heat Q_2 is rejected by the system at temp. T_2 . The change in entropy is S_2 . The change is negative.

4) From C_4 to C_1 , no heat is absorbed or rejected by the system, hence no change in entropy.

$\therefore$ Total change in entropy of a working substance during reversible cycle is $\frac{Q_1}{T_1} = \frac{Q_2}{T_2} \Rightarrow \frac{Q_1}{T_1} - \frac{Q_2}{T_2} = 0$

change in entropy of the system of source & sink is zero. Thus during reversible cycle, the entropy of the universe remains constant. i.e. $\Delta S_U = 0$.

* state & explain the principle of increase of entropy or explain entropy of universe in irreversible processes.

Efficiency of Carnot's cycle is

$$\eta = 1 - \frac{Q_2}{Q_1}$$

But, this efficiency is less than that of a reversible engine working between the same two temp's T_1 & T_2 . for which, efficiency

$$\eta = 1 - \frac{T_2}{T_1}$$

$$\therefore 1 - \frac{Q_2}{Q_1} < 1 - \frac{T_2}{T_1}$$

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Basic Thermodynamics B. K. VENKANNA, SWATI B. V., B. K. Venkanna, B. V. Swati, 2010-07 This book provides an in depth discussion of the principles of thermodynamics It focuses on engineering applications of theory and sound techniques for solving thermodynamic problems The book presents the fundamental concepts of thermodynamics and describes the theory of work and heat The text covers in detail the first law and the second law of thermodynamics with their applications It also explains the concepts of entropy and availability and irreversibility In addition the book presents thermodynamic properties of pure substances ideal gases and mixtures of ideal gases as well as real gases This book is designed for undergraduate students of mechanical engineering industrial and production engineering automobile engineering and aeronautical engineering for their courses in thermodynamics Key Features Presents the text in a simple and elegant manner to enable the students to grasp the essentials of the subject easily and quickly Covers all types of problems of various difficulty levels Includes more than 300 worked out examples and a large number of end of chapter exercises Provides solutions to several model question papers at the end of the book *Introduction to Applied Thermodynamics* R. M. Helsdon, 1965 **National Union Catalog**, 1978 Thermodynamics, Abridged William Duane Ennis, 1922 **Applied Thermodynamics** Henry Wilson Spangler, 1910 Chemical Abstracts Service Source Index American Chemical Society. Chemical Abstracts Service, 1990 The CAS Source Index CASSI Search Tool is an online resource intended to support researchers and librarians who need accurate bibliographic information This free resource can be used to quickly and easily look up or confirm publication titles and abbreviations as well as CODEN ISBN or ISSN codes The CASSI database contains a listing of publications indexed by Chemical Abstracts Service CAS since 1907 **The National Union Catalogs, 1963-**, 1964 **Applied Thermodynamics**, 2012 **Applied Thermodynamics Problems for Engineers** William John Peck, Alan James Richmond, 1962 **Applied Thermodynamics** Virgil Moring Faires, 1940

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