

WORK DONE IN THERMODYNAMICS

Total work done by gas

$$W = \int_{V_i}^{V_f} dW$$

If pressure is constant

$$W = P(V_f - V_i)$$

system → expand → volume increase

Work done by gas = +ve, Work done on gas = -ve

system → compression → volume decrease

Work done by gas = -ve, Work done on gas = +ve

$$W = \int P \cdot dV$$

Work done = Area under pressure – volume graph

Work done in cyclic process –

Work done is area inside the cycle.

Clock Wise → WD = +ve

Anticlockwise → WD = -ve

LAWS OF THERMODYNAMICS

Zeroth law of thermodynamics:

If A is in thermal equilibrium with B and also B is in thermal equilibrium with C then A will be in thermal equilibrium with C.

If, $T_A = T_C$ and also, $T_B = T_C$

Then, $T_A = T_B$

First Law of thermodynamics → [FLOT]

When heat is given to system then some part of this heat is used to do the work & some part increase its internal energy

$$Q = \Delta U + W$$

WD by system = +ve

WD on system = -ve

Temp of body ↓, $\Delta U = +ve$

Temp of body ↑ $\Delta U = -ve$

Heat is given to system = +ve

Heat is given by system = -ve

Second Law of Thermodynamics –

- Kelvin-plank statement:** It is impossible to construct an engine which can convert entire heat into useful work.
- Claussius's Statement:** In a self-acting machine it is not possible to transfer heat from colder body to hotter body without an aid of external energy.

Entropy: Degree of randomness or disorder

- Entropy is not a conserved quantity
- Entropy of universe always increase
- For an isolated system it is not possible for a process to decrease the entropy.

$$\Delta S = S_f - S_i = \int_{T_i}^{T_f} \frac{dQ}{T}$$

SPECIFIC HEAT CAPACITIES OF GASES

$$\frac{C_p}{C_v} = \gamma$$

$$C_p - C_v = R$$

$$C_v = \frac{R}{\gamma - 1}$$

$$C_p = \frac{\gamma R}{\gamma - 1}$$

$$\gamma_{mix} = \frac{n_1 C_{p1} + n_2 C_{p2}}{n_1 C_{v1} + n_2 C_{v2}}$$

$$\frac{n_1 + n_2}{\gamma_{mix} - 1} = \frac{n_1}{\gamma_1 - 1} + \frac{n_2}{\gamma_2 - 1}$$

$$C_{v(mix)} = \frac{n_1 C_{v1} + n_2 C_{v2}}{n_1 + n_2}$$

$$C_{p(mix)} = \frac{n_1 C_{p1} + n_2 C_{p2}}{n_1 + n_2}$$

FLOT FOR DIFFERENT PROCESSES

Isobaric Process: The physical change undergoing at constant pressure.

- $V \propto T$
- Slope of P-V curve = $\frac{dP}{dV} = 0$
- Specific heat

$$C_{p(mono)} = \frac{5R}{2}$$

$$C_{p(dia)} = \frac{7R}{2}$$

- Bulk modulus of elasticity, $B = -\frac{dP}{dV} \cdot V = 0$
- Work done : $W = P \cdot \Delta V = \text{Area under } P \cdot V \text{ curve}$

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