Name: \_

Date:

# STOICHIOMETRY PRACTICE WORKSHEET

1. Using the following equation:

Calculate how many grams of iodine are needed to prepare 28.6 grams of ICI by this reaction.

2. Using the following equation:

$$5 \text{ KNO}_2 + 2 \text{ KMnO}_4 + 3 \text{ H}_2 \text{SO}_4 \rightarrow 5 \text{ KNO}_3 + 2 \text{ MnSO}_4 + \text{ K}_2 \text{SO}_4 + 3 \text{ H}_2 \text{O}_4$$

How many moles and grams of  $\rm KMnO_4$  are needed for this reaction on 11.4 grams of  $\rm KNO_2?$ 

3. Using the following equation:

How many moles and grams of oxygen  $(O_2)$  are needed to react with 56.8 grams of ammonia by this reaction?

4. Using the following equation:

$$NaIO_3 + 6 HI \rightarrow 3 I_2 + NaI + 3 H_2O$$

Calculate the number of moles and the number of grams of iodine ( $l_2$ ) that can be made this way from 16.4 grams of NaIO<sub>3</sub>.

# **Basic Stoichiometry Practice Problems**

David M. Himmelblau, James B. Riggs

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