

Bonding in Metals (pages 36–39)

Metallic Bonding (page 37)

Key Concept: A metal crystal consists of positively charged metal ions embedded in a “sea” of valence atoms.

- The physical and chemical properties of metals can be explained by the structure of metal atoms and by the bonding between them.
- Most metals have one, two, or three valence electrons. Metals easily lose electrons. Metal ions have a positive electric charge.
- Solid metals are crystals. Metal atoms in the crystal are very close together. The atoms in a metal crystal are arranged in a specific way.
- In a metal crystal, the atoms exist as positive ions. The valence electrons lost from the ions freely drift around the ions in the crystal. The metal ions are held in place by metallic bonds. A **metallic bond** is the attraction between a positive metal ion and the electrons around it.

Answer the following questions. Use your textbook and the ideas above.

1. Circle the letter of each sentence that is true about metallic bonding.
 - a. Most metals have one, two, or three valence electrons.
 - b. Metals easily lose protons.
 - c. Metal ions have a positive electrical charge.

Bonding In Metals Guided Study Answers

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