

Biology 160 Exam 1 Study Guide

L1: Organization & Function of the Cell

5 Big Ideas:

1. **Evolution** – populations of organisms and cells, including their molecular components, that have changed over time through both selective and non-selective evolutionary processes.
2. **Structure & Function** – all living systems (cells, organisms, ecosystems, etc) are made of structural components whose arrangement determines the function of the systems.
3. **Information Flow & Storage** – information and signals are used and exchanged within and among cells and organisms to direct their functioning.
4. **Transformations of Energy and Matter** – all living things acquire, use, and release and cycle matter and energy for cellular / organismal functioning.
 - a. Life obeys the laws of physics and chemistry
 - b. All living things require input of energy and matter
 - c. Energy can be neither created nor destroyed, but can change form
5. **Systems** – living systems are interconnected, and they interact and influence each other on multiple levels.

L2: Chemical Elements & Chemical Bonds

- **Proton** – positively charged atomic particle located within the nucleus of an atom with a charge of 1+ and a mass of approximately 1 Da.
- **Electron** – negatively charged atomic particle located outside of the nucleus in the orbitals of the atom with a charge of 1- and a mass of approximately 0 Da.
- **Neutron** – a neutrally charged atomic particle located within the nucleus of the atom with no relative charge and a mass of approximately 1 Da.
- **Atomic Number** – the total # of protons in the nucleus of an atom.
- **Mass Number** – the total # of protons + neutrons in the nucleus of an atom.
- **Isotopes** – two atoms with the same # of protons but with DIFFERENT # of neutrons are considered to be isotopes of the same element.
 - o Electrons move around the nuclei in specific regions called orbitals & each orbital can hold up to 2 electrons.
- **Valence shell** – the outermost shell of electrons.
- **Valence** – the # of unpaired electrons in the valence shell of an atom as well as the number of bonds that atom is likely to form. (carbon = 4, nitrogen = 3)
 - o Unpaired electrons are relatively uncommon in nature & atoms that carry unpaired electrons are usually rather reactive.
 - o Unpaired electrons in the valence shell can form covalent bonds.
 - Ex – 2 hydrogen atoms (each with one unpaired electron) can form a covalent bond with 2 shared electrons.

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VM Jensen



Biology 160 Study Guide:

Reviewing **Biology 160 Study Guide**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Biology 160 Study Guide**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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Table of Contents Biology 160 Study Guide

1. Understanding the eBook Biology 160 Study Guide
 - The Rise of Digital Reading Biology 160 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Biology 160 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Biology 160 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Biology 160 Study Guide
 - Personalized Recommendations
 - Biology 160 Study Guide User Reviews and Ratings
 - Biology 160 Study Guide and Bestseller Lists

5. Accessing Biology 160 Study Guide Free and Paid eBooks
 - Biology 160 Study Guide Public Domain eBooks
 - Biology 160 Study Guide eBook Subscription Services
 - Biology 160 Study Guide Budget-Friendly Options
6. Navigating Biology 160 Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Biology 160 Study Guide Compatibility with Devices
 - Biology 160 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Biology 160 Study Guide
 - Highlighting and Note-Taking Biology 160 Study Guide
 - Interactive Elements Biology 160 Study Guide
8. Staying Engaged with Biology 160 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Biology 160 Study Guide
9. Balancing eBooks and Physical Books Biology 160 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Biology 160 Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Biology 160 Study Guide
 - Setting Reading Goals Biology 160 Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Biology 160 Study Guide
 - Fact-Checking eBook Content of Biology 160 Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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