

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

- **Crystalline Solids** have a regular, repeating arrangement of particles. They have sharp melting points and distinct X-ray diffraction patterns.
- **Amorphous Solids** lack a long-range order. They soften over a range of temperatures and do not have sharp X-ray diffraction patterns.
- **Unit Cells** are the smallest repeating units of a crystal lattice. Common types include simple cubic, body-centered cubic (BCC), and face-centered cubic (FCC).
- **Packing Efficiency** is the fraction of space occupied by atoms in a unit cell. FCC and HCP have the highest efficiency at approximately 74%.
- **Intermolecular Forces** (IMFs) determine the strength of the solid. Hydrogen bonding leads to higher melting points than other IMFs.
- **Band Theory** explains the electrical properties of solids. Metals have overlapping bands, while insulators have a large band gap.

Types of Chemical Calculations

Key to Answer on the Exam

- **Stoichiometry**: Calculations involving the relative amounts and masses of substances in a chemical reaction.
- **Thermodynamics**: Calculations involving energy changes, enthalpy, and entropy.
- **Equilibrium**: Calculations involving the concentrations of reactants and products at equilibrium.
- **Acid-Base**: Calculations involving pH, pKa, and the concentrations of acids and bases.
- **Solubility**: Calculations involving the concentrations of ions in a saturated solution.
- **Electrochemistry**: Calculations involving cell potentials, Gibbs free energy, and Faraday's laws.

Periodic Table & Trends

The periodic table is color-coded to show trends. Electronegativity increases from bottom-left to top-right. Ionization energy follows a similar trend. Atomic radius increases from top-left to bottom-right.

- **Electronegativity** is the ability of an atom to attract electrons in a bond. It increases across a period and decreases down a group.
- **Ionization Energy** is the energy required to remove an electron from a gaseous atom. It increases across a period and decreases down a group.
- **Atomic Radius** is the distance from the nucleus to the outermost electrons. It increases down a group and across a period.
- **Electron Affinity** is the energy change when an electron is added to a gaseous atom. It is most negative for halogens.

Electronegativity Table

Electronegativity Trends:

- **Increases** from bottom-left to top-right.
- **Fluorine (F)** is the most electronegative element (4.0).
- **Francium (Fr)** is the least electronegative element (0.7).

The table shows electronegativity values for elements 1 through 118. Values range from 0.7 for Francium to 4.0 for Fluorine.

Thermodynamics

- **First Law of Thermodynamics**: Energy is conserved. $\Delta E = q + w$
- **Enthalpy (H)**: A state function. $\Delta H = H_{\text{products}} - H_{\text{reactants}}$
- **Calorimetry**: The study of heat changes. Bomb calorimetry measures ΔU , while coffee cup calorimetry measures ΔH .
- **Hess's Law**: The total enthalpy change for a reaction is independent of the path taken.

Equilibrium & Solubility

- **Equilibrium Constant (K)**: A measure of the extent of a reaction. $K = \frac{[\text{products}]}{[\text{reactants}]}$
- **Le Chatelier's Principle**: A system at equilibrium will shift to counteract a change in conditions.
- **Solubility Product (K_{sp})**: The product of the concentrations of ions in a saturated solution.

Compound	K _{sp}
AgCl	1.8 × 10 ⁻¹⁰
PbCl ₂	1.7 × 10 ⁻⁵
CaF ₂	3.9 × 10 ⁻¹¹
Fe(OH) ₃	4 × 10 ⁻³⁸
Al(OH) ₃	3 × 10 ⁻³⁴
CuS	8.5 × 10 ⁻⁴⁵
SnS	1 × 10 ⁻²⁵
As ₂ S ₃	6.8 × 10 ⁻³³
Sb ₂ S ₃	1 × 10 ⁻²³
Bi ₂ S ₃	1 × 10 ⁻⁷²
FeS	8 × 10 ⁻¹⁹
ZnS	2 × 10 ⁻²²
Al(OH) ₃	3 × 10 ⁻³⁴
Cr(OH) ₃	6.3 × 10 ⁻³¹
Mn(OH) ₂	2 × 10 ⁻¹³
Ni(OH) ₂	2 × 10 ⁻¹⁶
Co(OH) ₂	2 × 10 ⁻¹⁶
Cu(OH) ₂	2 × 10 ⁻²⁰
Pb(OH) ₂	1.2 × 10 ⁻¹⁵
Ba(OH) ₂	5 × 10 ⁻⁶
Ca(OH) ₂	5.5 × 10 ⁻⁶
NaOH	> 10 ¹⁴
KOH	> 10 ¹⁴



Acids & Bases

- **Arrhenius Definition**: Acids produce H⁺ ions in water; bases produce OH⁻ ions.
- **Brønsted-Lowry Definition**: Acids are proton donors; bases are proton acceptors.
- **Conjugate Acid-Base Pairs**: Differ by one proton. $\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-$
- **pH and pKa**: $\text{pH} = -\log[\text{H}^+]$; $\text{pKa} = -\log K_a$
- **Titration Curves**: Graphs of pH vs. volume of titrant. The equivalence point is where moles of acid = moles of base.

Redox Reactions

- **Oxidation**: Loss of electrons. **Reduction**: Gain of electrons.
- **Oxidation States**: Assigned to atoms to track electron flow. Rules include: elements are 0, monatomic ions are their charge, etc.
- **Half-Reactions**: Separate oxidation and reduction processes.
- **Cell Potentials**: $E_{\text{cell}} = E_{\text{cathode}} - E_{\text{anode}}$
- **Gibbs Free Energy**: $\Delta G = -nFE_{\text{cell}}$
- **Faraday's Laws**: Relate the amount of substance to the amount of electricity.

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