

Activity 10

Classifying reactions

Can you classify these?

Type of chemical reactions and their descriptions



Introduction: Chemical reactions show how **compounds** and **elements** react with one another. An element is an substance consisting of only one of a kind. Such as the elements of hydrogen, oxygen and fluorine. A compound is a substance made of more than one kind of atoms. Such as water, which is made up of hydrogen.

Question: How are chemical reactions classified?

1. Study these chemical reactions and try to classify each one of four types. Using the chemical equations in the boxes on the right, match the following definitions to the type of reaction.

☒ The reaction involves two or more elements.

a. **Redox**

☒ A reaction involving with oxygen as product which burns and heats.

a. **Exothermic**

☒ Two or more compounds combine to form one product.

c. **Single replacement**

☒ Two compounds react to form two different compounds.

b. **Double replacement**

☒ A compound reacts with an element to form a new compound and a different element.

d. **Combustion**

2. Study these examples of the chemical reactions below. Some reactions may already be classified in the space on the right. Identify the reaction as a synthesis, decomposition, single replacement, or double replacement reaction.

a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ **Synthesis**

Double

b. $\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$ **Decomposition**

Single

c. $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$ **Single replacement**

d. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ **Double replacement**

e. $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$ **Single replacement**

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S Ashworth

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