

Stoichiometry Lab - **ANSWERS****PRE-LAB QUESTIONS:**

1. Calculate the molar mass of sodium bicarbonate (NaHCO_3). 84.003 g/mol
 2. Calculate the molar mass of acetic acid (CH_3COOH). 60.052 g/mol
 3. Calculate the molar mass of carbon dioxide (CO_2). 44.003 g/mol
 4. Vinegar is generally 4-5% acetic acid dissolved in distilled water. That means that a 5% vinegar solution contains 0.05 mol of acetic acid per liter of solution. How many moles of acetic acid are present in 30.00 mL?
- $$\frac{0.05 \text{ mol CH}_3\text{COOH}}{1 \text{ liter vinegar}} \times \frac{0.03 \text{ liter CH}_3\text{COOH}}{1 \text{ liter vinegar}} = 0.0015 \text{ mol CH}_3\text{COOH in 30.0 mL of vinegar}$$
5. When baking soda (sodium bicarbonate) reacts with vinegar (acetic acid), they form sodium acetate, water, and carbon dioxide gas. Balance the chemical equation for this reaction.



6. A student measures out 5.00 g of baking soda to react with 50.00 mL of vinegar (which contains 0.042 moles of acetic acid).

a) Determine the limiting reagent, excess reagent, and theoretical yield (in grams) of carbon dioxide. Show your work.

$$5.00 \text{ g baking soda} \times \frac{1 \text{ mol baking soda}}{84.003 \text{ g baking soda}} = \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol baking soda}} \times \frac{44.003 \text{ g carbon dioxide}}{1 \text{ mol carbon dioxide}} = 2.62 \text{ g carbon dioxide}$$

$$0.042 \text{ mole acetic acid} \times \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol acetic acid}} \times \frac{44.003 \text{ g carbon dioxide}}{1 \text{ mol carbon dioxide}} = 1.85 \text{ g carbon dioxide}$$

*** Since acetic acid produced the least amount of product, the acetic acid in vinegar is the **limiting reagent** and the baking soda is the **excess reagent**.

*** The theoretical yield of carbon dioxide is 1.85 g.

- b) This reaction will produce carbon dioxide bubbles that escape. After the reaction, the mass of the reaction decreased by 1.50 g. This is the actual mass of carbon dioxide that was produced from the reaction. Calculate the percentage yield of carbon dioxide. Show your work.

$$\text{Percent Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{Percent Yield} = \frac{1.50 \text{ g}}{1.85 \text{ g}} \times 100 = 79.4\% \text{ yield}$$

Baking Soda And Vinegar Stoichiometry Lab Answers

Connie Lynn Bannick Kemner



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