



Calcium oxalate crystals contribute to the formation of kidney stones.

Experiment 8

Limiting Reactant

- To determine the limiting reactant in a mixture of two soluble salts
- To determine the percent composition of each substance in a salt mixture

OBJECTIVES

The following techniques are used in the Experimental Procedure:



TECHNIQUES

Percent composition: the mass ratio of a component of a mixture or compound to the total mass of the sample times 100.

Two factors affect the yield of products in a chemical reaction: (1) the amounts (moles) of starting materials (reactants) and (2) the **percent yield** of the reaction. Many experimental conditions, for example, temperature and pressure, can be adjusted to increase the yield of a desired product in a chemical reaction, but because chemicals react according to fixed molar ratios (**stoichiometrically**), only a limited amount of product can form from measured amounts of starting materials. The reactant determining the amount of product generated in a chemical reaction is called the **limiting reactant** in the chemical system.

To better understand the concept of the limiting reactant, let us look at the reaction under investigation in this experiment, the reaction of calcium chloride dihydrate, $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$, and potassium oxalate monohydrate, $\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$, in an aqueous solution.



Calcium oxalate monohydrate, $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$, the product, is an insoluble compound, but is found naturally in a number of diverse locations. It is found in plants, such as rhubarb leaves, eggs, and in small amounts spinach, and is the cause of most kidney stones. In small doses, it causes a severe reaction to the lining of the digestive tract. However, the handling of calcium oxalate in the laboratory is safe, so long as it is not transferred to the mouth.

For the reaction system in this experiment, both the calcium chloride and potassium oxalate are soluble salts, but the calcium oxalate is insoluble. The **ionic equation** for the reaction is



INTRODUCTION

Percent yield

$$\left(\frac{\text{actual yield}}{\text{theoretical yield}} \right) \times 100$$

Stoichiometrically: by a study of a chemical reaction using a balanced equation.

Ionic equation: A chemical equation that presents ionic compounds in the form in which they exist in aqueous solution. (See Experiment 8.)

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M Tight



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