

Section 3.1

P. 173 Review Questions #1-14

1. *Sample answer:* People call Thomson's model the "plum pudding" model because the electrons were embedded in a positively charged sphere like plums in a pudding. Plum pudding was common in Thomson's time but now, many people haven't even heard of it. So we could maybe think of the electrons as blueberries in blueberry muffins.
2. **a.** In a chemical reaction, the mass of products equals the mass of reactants. According to Dalton's theory, atoms are not destroyed or created in chemical reactions—instead they are rearranged. Therefore, the mass doesn't change.
b. Dalton's theory could not explain why atoms combined in certain ratios.
3. The discovery of the neutron solved the problem. The neutron adds mass to the nucleus without adding any charge.



5. In order for the foil to be penetrated by alpha particles, it had to be extremely thin. Because gold is a highly malleable metal, it could be used to make an extremely thin foil.
6. Thomson and Rutherford: atoms contain positive and negative charges, contains electrons; Thomson: electrons distributed throughout sphere, positive charge spread throughout sphere; Rutherford: positive charge in tiny, dense nucleus; electrons orbit the nucleus.
7. Rutherford's electrons are free to orbit anywhere in the volume of space around the nucleus; Bohr's have only specific allowable energy levels. This change provides a model that can explain emission and absorption spectra.

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Kokuritsu Kokkai Toshokan (Japan)



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