



Calcium oxalate crystals contribute to the formation of kidney stones.

Experiment 8

Limiting Reactant

- To determine the limiting reactant in a mixture of two soluble salts
- To determine the percent composition of each substance in a salt mixture

OBJECTIVES

The following techniques are used in the Experimental Procedure:



TECHNIQUES

Percent composition: the mass ratio of a component of a mixture or compound to the total mass of the sample times 100.

Two factors affect the yield of products in a chemical reaction: (1) the amounts (moles) of starting materials (reactants) and (2) the **percent yield** of the reaction. Many experimental conditions, for example, temperature and pressure, can be adjusted to increase the yield of a desired product in a chemical reaction, but because chemicals react according to fixed molar ratios (**stoichiometrically**), only a limited amount of product can form from measured amounts of starting materials. The reactant determining the amount of product generated in a chemical reaction is called the **limiting reactant** in the chemical system.

To better understand the concept of the limiting reactant, let us look at the reaction under investigation in this experiment, the reaction of calcium chloride dihydrate, $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$, and potassium oxalate monohydrate, $\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$, in an aqueous solution.



Calcium oxalate monohydrate, $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$, the product, is an insoluble compound, but is found naturally in a number of diverse locations. It is found in plants, such as rhubarb leaves, eggs, and in small amounts spinach, and is the cause of most kidney stones. In small doses, it causes a severe reaction to the lining of the digestive tract. However, the handling of calcium oxalate in the laboratory is safe, so long as it is not transferred to the mouth.

For the reaction system in this experiment, both the calcium chloride and potassium oxalate are soluble salts, but the calcium oxalate is insoluble. The **ionic equation** for the reaction is



INTRODUCTION

Percent yield

$$\left(\frac{\text{actual yield}}{\text{theoretical yield}} \right) \times 100$$

Stoichiometrically: by a study of a chemical reaction using a balanced equation.

Ionic equation: A chemical equation that presents ionic compounds in the form in which they exist in aqueous solution. (See Experiment 8.)

Beran Lab Manual Experiment 8

RS Peters



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Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16

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Laboratory Manual Jo A. Beran, 1990

Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance

El-Hi Textbooks & Serials in Print, 2000, 2000

Chemistry in the Laboratory Jo A. Beran, 1996 This updated and revised chemistry manual provides a well rounded understanding of concepts in the general chemistry laboratory Utilising visual aids experiments and equipment are explained and results and their pertinence discussed

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