

Задача 11.
Слабые и сред. кислоты

① Дано:

$$C_{\text{HNO}_2} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-4}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-5} = 4,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HNO}_2}}$$

$$= \sqrt{7,9 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-5} \text{ моль/л}$$

$$= 4,55$$

② Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}^+} = C_{\text{OH}^-}$$

$$K_{\text{Pr}} = [C_{\text{H}^+}] [C_{\text{OH}^-}]$$

$$= 6,3 \cdot 10^{-5} = 4,5^2$$

5 - концентрация ионов

$$5 = \sqrt{\frac{K_{\text{Pr}}}{C_{\text{H}^+}}} = \sqrt{\frac{6,3 \cdot 10^{-5}}{C_{\text{H}^+}}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[C_{\text{H}^+}] = \frac{5}{2} = 2,5 \cdot 10^{-2} \text{ моль/л}$$

$$[C_{\text{H}^+}] = \frac{10^{-14}}{[C_{\text{OH}^-}]} = \frac{10^{-14}}{2,5 \cdot 10^{-2}} = 4,25 \cdot 10^{-13}$$

$$pH = -\lg [C_{\text{H}^+}] = -\lg 4,25 \cdot 10^{-13} = 12,37$$

③ Дано:

$$V_{\text{HNO}_2} = 10 \text{ мл}$$

$$C_{\text{HNO}_2} = 0,1 \text{ м}$$

$$V_{\text{PrOH}_2} = 10 \text{ мл}$$

$$C_{\text{PrOH}_2} = 0,1 \text{ м}$$

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

Решение:

полн. смешивание:

$$C_{\text{HNO}_2} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

Для диссоциирующей кислоты:

$$C_{\text{H}^+} = C_{\text{HNO}_2} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = C_{\text{PrOH}_2} = 0,05 \text{ моль/л}$$

$$[C_{\text{Pr}^+}] [C_{\text{Cl}^-}] = 6,05 \cdot 10^{-5} = 10^{-12}$$



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