

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

- **Crystalline Solids:** Atoms, ions, or molecules are arranged in a regular, repeating three-dimensional pattern. They have sharp melting points and distinct X-ray diffraction patterns.
- **Amorphous Solids:** Atoms, ions, or molecules are arranged in a disordered, non-repeating pattern. They do not have sharp melting points and do not show distinct X-ray diffraction patterns.
- **Intermolecular Forces:** The forces between particles in a solid determine its properties. Stronger forces lead to higher melting points and greater hardness.
- **Types of Solids:**
 - Covalent Network Solids:** Atoms are bonded together in a continuous network of covalent bonds. Examples include diamond, silicon, and quartz.
 - Ionic Solids:** Consist of alternating positive and negative ions held together by strong electrostatic forces. Examples include NaCl and CaF₂.
 - Molecular Solids:** Consist of discrete molecules held together by intermolecular forces. Examples include ice, sugar, and iodine.
 - Metallic Solids:** Consist of metal cations in a sea of delocalized electrons. They are good conductors of electricity and heat.

Types of Chemical Calculations Allowed for Answer on the Exam

- **Stoichiometry:** Calculations involving the relative masses and numbers of atoms, molecules, or ions in a chemical reaction.
- **Thermochemistry:** Calculations involving the heat changes that accompany chemical reactions.
- **Equilibrium:** Calculations involving the concentrations of reactants and products at equilibrium.
- **Acid-Base Chemistry:** Calculations involving the pH of solutions and the equilibrium constants for acid-base reactions.
- **Electrochemistry:** Calculations involving the relationship between electrical energy and chemical reactions.

Periodic Table & Trends

The periodic table is color-coded to show trends. Electronegativity increases from bottom-left to top-right. Ionization energy increases from bottom-left to top-right. Atomic radius increases from top-left to bottom-right.

- **Electronegativity:** The ability of an atom to attract electrons in a chemical bond. It increases across a period and decreases down a group.
- **Ionization Energy:** The energy required to remove an electron from a gaseous atom. It increases across a period and decreases down a group.
- **Atomic Radius:** The distance from the nucleus to the outer edge of the electron cloud. It increases down a group and across a period.

Electronegativity Table

Electronegativity (Pauling Scale)

Electronegativity is a measure of an atom's ability to attract and hold onto electrons in a chemical bond. It is a relative property, with fluorine being the most electronegative element (4.0) and cesium being the least (0.7).

Electronegativity values generally increase from left to right across a period and from bottom to top within a group. This trend is due to the increasing effective nuclear charge and decreasing atomic radius, which result in a stronger attraction for electrons.

Electronegativity differences between atoms can be used to predict the type of bond that will form between them:

- Nonpolar covalent bond:** Formed between atoms with a small electronegativity difference (usually less than 0.5).
- Polar covalent bond:** Formed between atoms with a moderate electronegativity difference (usually between 0.5 and 1.7).
- Ionic bond:** Formed between atoms with a large electronegativity difference (usually greater than 1.7).

The periodic table below shows electronegativity values for various elements, color-coded by groups:

- Blue: Group 1 (Alkali Metals)
- Green: Groups 2-10 (Transition Metals)
- Yellow: Groups 11-18 (Post-Transition Metals and Nonmetals)
- Red: Group 17 (Halogens)
- White: Group 18 (Noble Gases)

Electronegativity values (Pauling Scale) for various elements:

Element	Electronegativity
Li	0.9
Be	1.5
B	2.0
C	2.5
N	3.0
O	3.5
F	4.0
Ne	-
Na	0.9
Mg	1.2
Al	1.5
Si	1.8
P	2.1
S	2.5
Cl	3.0
Ar	-
K	0.8
Ca	1.0
Sc	1.3
Ti	1.5
V	1.6
Cr	1.6
Mn	1.5
Fe	1.8
Co	1.8
Ni	1.9
Cu	1.9
Zn	1.6
Ga	1.6
Ge	1.8
As	2.0
Se	2.4
Br	2.8
Kr	-
Rb	0.8
Sr	1.0
Y	1.3
Zr	1.4
Nb	1.6
Mo	1.7
Tc	1.7
Ru	1.8
Rh	1.9
Pd	2.2
Ag	1.9
Cd	1.7
In	1.7
Sn	1.8
Sb	2.0
Te	2.1
I	2.5
Xe	-
Ba	0.9
La	1.0
Ce	1.1
Pr	1.1
Nd	1.1
Pm	1.1
Sm	1.0
Eu	1.0
Gd	1.0
Tb	1.0
Dy	1.0
Ho	1.0
Er	1.0
Tm	1.0
Yb	1.0
Lu	1.0
Hf	1.3
Ta	1.5
W	1.7
Re	1.7
Os	1.8
Ir	1.9
Pt	2.2
Au	2.4
Hg	1.9
Tl	1.8
Pb	1.8
Bi	1.9
Po	2.0
At	2.2
Fr	0.7
Ra	0.9
Ac	1.0
Th	1.3
Pa	1.3
U	1.3
Np	1.3
Pu	1.3
Am	1.3
Cm	1.3
Bk	1.3
Cf	1.3
Es	1.3
Fm	1.3
Md	1.3
No	1.3
Lr	1.3

Ap Chemistry Quick Study Academic Wallchart Common

Brendan G. Carr



Ap Chemistry Quick Study Academic Wallchart Common:

The Top Books of the Year Ap Chemistry Quick Study Academic Wallchart Common The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous engrossing novels captivating the hearts of readers worldwide. Lets delve into the realm of bestselling books, exploring the fascinating narratives that have enthralled audiences this year. Ap Chemistry Quick Study Academic Wallchart Common : Colleen Hoover's "It Ends with Us" This touching tale of love, loss, and resilience has gripped readers with its raw and emotional exploration of domestic abuse. Hoover masterfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can triumph. Uncover the Best : Taylor Jenkins Reid's "The Seven Husbands of Evelyn Hugo" This captivating historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reid's absorbing storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Ap Chemistry Quick Study Academic Wallchart Common : Delia Owens' "Where the Crawdads Sing" This mesmerizing coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens spins a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These top-selling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a brilliant and suspenseful novel that will keep you wondering until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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